

Article

Catalytic Upgrading of Residual Fat Pyrolysis Vapors over Activated Carbon Pellets into Hydrocarbons-Like Fuels in A Two-Stage Reactor: Analysis of Chemical Kinetics and Physical-Chemistry Properties

Lucas Pinto Bernar¹, Caio Campos Ferreira¹, Augusto Fernando de Freitas Costa¹, Haroldo Jorge da Silva Ribeiro¹, Wenderson Gomes dos Santos¹, Lia Martins Pareira¹, Anderson Mathias Pereira¹, Nathalia Lobato Moraes², Fernanda Paula da Costa Assunção³, Sílvia Alex Pereira da Mota⁴, Douglas Alberto Rocha de Castro⁵, Marcelo Costa Santos¹, Neyson Martins Mendonça², Sergio Duvoisin Jr.⁶, Luiz Eduardo Pizarro Borges⁷ and Nélío Teixeira Machado^{1,2*}

¹ Graduate Program of Natural Resources Engineering of Amazon, Campus Profissional-UFPA, Universidade Federal do Pará, Rua Augusto Corrêa N° 1, Belém 66075-110, Brazil; lucas.bernar7@gmail.com (L.P.B.); caiocf7@hotmail.com (C.C.F.); affreitas@ufpa.br (A.F.d.F.C.); ribeiroengq@hotmail.com (H.J.d.S.R.); liapereira@ufpa.br (L.M.P.); wendersoneq@hotmail.com (W.G.d.S.); marcelo.santos@ufpa.edu.br (M.C.S.)

² Faculty of Sanitary and Environmental Engineering, Campus Profissional-UFPA, Universidade Federal do Pará, Rua Corrêa N° 1, Belém 66075-900, Brazil; nathalia.morras@icb.ufpa.br (N.L.M.); yasmincosta.eng@hotmail.com (Y.S.C.); neysonmm@ufpa.br (N.M.M.)

³ Graduate Program of Civil Engineering, Campus Profissional-UFPA, Universidade Federal do Pará, Rua Augusto Corrêa N° 1, Belém 66075-110, Brazil; fernanda.assuncao.itec@gmail.com (F.P.d.C.A.)

⁴ Graduate Program of Chemistry, Universidade Federal do Sul e Sudeste do Pará, Folha 31, Quadra 7, Lote Especial - Nova Marabá, CEP: 68.507.590, Marabá/PA, Brasil, silviomota@unifesspa.edu.br (S.A.P.d.M.)

⁵ Centro Universitário Luterano de Manaus – CEULM/ULBRA, Avenida Carlos Drummond de Andrade N°. 1460, Manaus 69077-730, Brazil; douglascastro87@hotmail.com (D.A.R.d.C.)

⁶ Faculty of Chemical Engineering, Universidade do Estado do Amazonas-UEA, Avenida Darcy Vargas N°. 1200, Manaus 69050-020, Brazil; sjunior@uea.edu.br (S.D.Jr.)

⁷ Laboratory of Catalyst Preparation and Catalytic Cracking, Section of Chemical Engineering, Instituto Militar de Engenharia-IME, Praça General Tibúrcio N°. 80, Rio de Janeiro 22290-270, Brazil; luiz@ime.eb.br (L.E.P.B.)

* Correspondence: machado@ufpa.br; Tel.: +55-91-984-620-325

Abstract: This work aims to investigate the influence of reaction time and catalyst-to-residual fat ratio by catalytic upgrading from pyrolysis vapors of residual fat at 400 °C and 1.0 atmosphere, on the yields of reaction products, physicochemical properties (density, kinematic viscosity, and acid value) and chemical composition of bio-oils, over a catalyst fixed bed reactor of activated carbon pellets impregnated with 10.0 M NaOH, in semi pilot scale. The experiments were carried out at 400 °C and 1.0 atmosphere, using a process schema consisting of a thermal cracking reactor of 2.0 L coupled to a catalyst fixed bed reactor of 53 mL, without catalyst and using 5.0, 7.5, and 10.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, in batch mode. Samples of liquid phase products were withdrawn during the course of reaction at 50, 60, 70, 80, 90, 100 and 120 minutes in order to investigate the process kinetics. The physicochemical properties (density, kinematic viscosity, and acid value) of bio-oils were determined by official methods. The chemical composition of bio-oils determined by GC-MS. The thermal catalytic cracking of residual fat show bio-oils yields from 55.55 to 30.22 (wt.%), aqueous phase yields between 2.83 and 3.19 (wt.%), solid phase yields between 13.56 and 9.75 (wt.%), and gas yields from 27.89 to 55.60 (wt.%). The yields of bio-oil decreases from 74.41 to 30.22% (wt.) with increasing catalyst-to-Tallow kernel oil ratio, while that of gaseous phase increases from 12.87 to 55.60% (wt.). For all the thermal and thermal catalytic cracking experiments, the density, kinematic viscosity, and acid value of bio-oils

decreases as the reaction time increases varying from 0.9266 to 0.8220 g/cm³, 8.10 to 2.24 mm²/s, and 144.14 to 2.37 mg KOH/g. The GC-MS of liquid reaction products identified the presence of hydrocarbons (alkanes, alkenes, ring-containing alkanes, ring-containing alkenes, and aromatics) and oxygenates (carboxylic acids, ketones, esters, alcohols, and aldehydes). For all the pyrolysis and catalytic cracking experiments, the hydrocarbon selectivity in bio-oil increases with increasing reaction time, while those of oxygenates decrease, reaching concentrations of hydrocarbons up to 95.35% (area.). The best results for the physicochemical properties density, kinematic viscosity, and acid value were 0.8220 g/cm³, 3.03 mm²/s, and 2.37 mg KOH/g, respectively, with a maximum hydrocarbon concentration of 97.194% (area.) and 2.806% ketones (area.) were obtained at 400 °C and 1.0 atmosphere, 80 minutes, without catalyst. For the catalytic cracking experiments, the maximum hydrocarbon content of 75.763% (area.) and 17.041% (area.) carboxylic acids, 4.702% (area.) ketones (area.), and 2.494% (area.) non-identified oxygenates was obtained at 400 °C and 1.0 atmosphere, 90 minutes, using a catalyst fixed bed reactor, with 10.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH as catalyst.

Keywords: residual fat; activated carbon pellets; chemical activation; thermal catalytic cracking; catalyst bed reactor; liquid hydrocarbons

1. Introduction

In the last years, catalytic upgrading of lipid-base pyrolysis vapors over catalyst fixed bed has been applied as an alternative process to deoxygenate bio-oils volatiles and produce liquid mixtures rich in hydrocarbons with superior transport and physicochemical properties, and the literature reports several studies on the subject [1-13].

Most studies on catalytic upgrade focused on the pyrolysis vapors of lipid-base materials including *microalgae* such as *Chlorella vulgaris* [1], *Chromolaena odorata* [2], *Jatropha* waste [3-6], *soap stock* [2,7-8], *waste cooking oil* [9-10], *cottonseed oil dregs* [11], *tallow kernel oil* [12], and *castor seed oil* [13].

The pyrolysis vapors of lipid-base materials chemically upgraded over catalyst supports including Ni/Y-zeolite [1], HZSM-5 [2, 6-8, 10], metal oxide (Ce, Pd, Ru, Ni) impregnated activated carbon [3], oxide-base catalysts [4-5, 13], including aluminum (Al₂O₃, CeO₂/Al₂O₃, Pd/CeO₂/Al₂O₃, Ru/CeO₂/Al₂O₃, Ni/CeO₂/Al₂O₃), zirconia (ZrO₂, CeO₂/ZrO₂, Pd/CeO₂/ZrO₂, Ru/CeO₂/ZrO₂, Ni/CeO₂/ZrO₂) and titania based catalysts (TiO₂-R, TiO₂-A, CeO₂/TiO₂-R, CeO₂/TiO₂-A, ZrO₂/TiO₂, Pd/CeO₂/TiO₂-R, Ru/CeO₂/TiO₂-R, Ni/CeO₂/TiO₂-R, Pd/CeO₂/TiO₂-A, Ru/CeO₂/TiO₂-A, Ni/CeO₂/TiO₂-A) [4], CaO [5, 13], ZnO [13], and kaolin [13], MCM-41 [8, 12], HZSM-5/MCM-41 [8], metal oxide (Ca, Ce, Zr, Ni, Co) impregnated HZSM-5 (CaO/HZSM-5, CeO₂/HZSM-5, ZrO₂/HZSM-5, NiO/HZSM-5, CoO/HZSM-5) [9], metal impregnated (Co, Ni) HZSM-5 (Co/HZSM-5, Ni/HZSM-5), and Mo-Ni supported on γ -Al₂O₃ (α -Al₂O₃, γ -Al₂O₃, Mo-Co/ γ -Al₂O₃, and Mo-Ni/ γ -Al₂O₃).

All the studies on catalytic upgrade of pyrolysis vapors from lipid-base materials focused on deoxygenation of bio-oil [1-13], but emphasis has also been given on the conversion of BTEXN [3, 6, 9], as well as reaction mechanism/pathway [3, 10-12]. The catalytic upgrade of pyrolysis volatiles has been carried out by flash pyrolysis/analytical pyrolysis (Py-GC/MS) [3-4, 6], as well as by vacuum pyrolysis [7, 9, 11-12], in drop-tube/downdraft reactors [2, 7-9, 12], fixed bed reactors [1-2, 5, 7-9, 11-13], and fluidized bed reactors [10]. The catalytic upgrade of pyrolysis vapors were performed in analytical [3-4, 6], micro [1, 11], and laboratory scales [2, 5, 7-10, 12-13]. The catalytic upgrade

processes operated in batch [1, 3-4, 6, 13], and continuous mode [2, 5, 7-12], and only a few studies operated as a two-stage reactor, that is, using a process schema consisting of a thermal cracking reactor coupled to a catalyst fixed bed reactor [2, 7-9, 12-13].

The reaction products by the catalytic upgrade of pyrolysis bio-oils from *lipid-base materials* [1-13], include gaseous and liquid fuels, water, aqueous acid phase, and coke [2, 6-10, 12-13].

The investigated physicochemical properties by the catalytic upgrade of pyrolysis bio-oils includes kinematic viscosity [13], high heating value [4, 7, 8, 11, 13], cloud point [13], acidity [11], water content [1, 4, 7, 8, 10, 13], and pH [13].

The pyrolysis bio-oil are composed by alkanes, alkenes, ring-containing alkanes, ring-containing alkenes, cyclo-alkanes, cyclo-alkenes, and aromatics [1-13], and oxygenates including carboxylic acids, aldehydes, ketones, fatty alcohols, phenols, amines, amides, ethers, and esters [1-13].

Beyond the operating mode (batch, continuous), type of pyrolysis process (flash pyrolysis/analytical pyrolysis, flash pyrolysis, and vacuum pyrolysis), type of reactors (drop-tube reactors, downdraft, fixed bed reactors, and fluidized bed reactors), as well as process schema (two-stage reactor), other process parameters that may affect the yields and chemical composition of bio-oil by catalytic upgrade of pyrolysis vapors are temperature [1-2, 7-8, 12], catalyst-to-biomass [1-4, 6-8, 12-13], characteristics of feed materials [1-13], weight hour space velocity [7], and the process scale including analytical [3-4, 6], micro [1, 11], and laboratory [2, 5, 7-10, 12-13].

Despite some studies focusing the effect of catalyst-to-biomass ratio on the yield and chemical composition of bio-oil by catalytic upgrade of pyrolysis vapors in micro [1], and laboratory scales [2, 7-8, 12-13], *until the moment, no systematic study has investigated the effect of catalyst-to-biomass ratio on the yield, chemical composition and physicochemical properties (density, kinematic viscosity, and acidity) of bio-oil pyrolysis in semi-pilot scale, using a two-stage reactor schema, as well as the influence reaction time on hydrocarbons and oxygenates composition and physicochemical properties of bio-oil pyrolysis using activated carbon pellets impregnated with 10.0 M NaOH.* The effect of catalyst-to-biomass ratio on the yield and chemical composition of bio-oil pyrolysis by catalytic upgrade of pyrolysis vapors [6, 12-13], summarized as follows.

Vichaphund *et. al.* [6], investigated the influence of catalyst-to-biomass ratio on the yield and chemical composition of bio-oil organic phase by the catalytic upgrading of *Jatropha* waste pyrolysis vapors over Co and Ni impregnated HZSM-5, nanocrystalline HZSM-5, using flash pyrolysis/analytical pyrolysis (Py-CG/MS). The experiments were carried out at 500 °C, 30 seconds, using HZSM-5, Co/HZSM-5, and Ni/HZSM-5, and catalyst-to-biomass ratios of 1.0, 5.0, and 10.0. The experiments show for all the catalysts the yield of bio-oil organic phase decreases with increasing catalyst-to-biomass ratios. The GC-MS analysis show that hydrocarbons content (aliphatic, aromatics) increases drastically with increasing catalyst-to-biomass ratios, varying between 30.43 and 96.3% (area.%) for HZSM-5, between 31.96 and 95.33% (area.%) for Co/HZSM-5, and between 34.76 and 97.55% (area.%) for Ni/HZSM-5. The highest yield of bio-oil organic phase was obtained with Ni/HZSM-5, catalyst-to-biomass ratio of 1:1, being approximately to 53.0% (wt.). The Py-GC-MS analysis identified hydrocarbons (aliphatic and aromatic) and oxygenates (esters, ethers, sugars, phenols, ketones, alcohols, and carboxylic acids).

Yu *et. al.* [13], investigated the effect of catalyst-to-Tallow kernel oil ratio and process temperature on the yield and chemical composition of bio-oil by the catalytic upgrading of vapors from Tallow kernel oil pyrolysis over SiC foam-MCM41, using quartz annular tube placed inside a series of 02 (two) microwave ovens, whereas the vapors produced by pyrolysis in the first oven flows into the

second microwave oven containing the fixed bed of catalysts. The pyrolysis reactor temperature were 450, 500, 550, 600, and 650 °C, while the catalyst bed reactor temperature were 200, 250, 300, 350 and 400 °C, and catalyst-to-Tallow kernel oil of 1:2, 2:1, 3:2, 1:1, and 2:3. For the experiments maintaining constant the catalyst-to-tallow kernel oil ratio of 1:2, catalyst bed temperature of 350 °C, and varying the pyrolysis temperature between 450 and 650 °C, the yield of bio-oil decreases as the temperature increases. In addition, the concentration of hydrocarbons increases from 53.18 to 93.20% (area.) between 450 and 600 °C, while that of oxygenates decrease from 46.82 to 6.80% (area.). For the experiments maintaining constant the catalyst-to-tallow kernel oil ratio of 1:2, pyrolysis temperature of 600 °C, and varying the catalyst bed temperature between 200 and 400 °C, the yield of bio-oil increases as the temperature increases. In addition, the concentration of hydrocarbons increases from slightly from 89.14 to 93.20% (area.) between 200 and 350 °C, while that of oxygenates decrease from 10.76 to 6.80% (area.). For the experiments maintaining constant the pyrolysis temperature of 600 °C, the catalyst bed temperature of 350 °C, and varying the catalyst-to-tallow kernel oil ratio between 2:1 and 2:3, the yield of bio-oil decreases as the catalyst-to-Tallow kernel oil ratio increases. It was also observed that catalyst-to-Tallow kernel oil ratio had no effect on the hydrocarbon selectivity. The highest bio-oil yield of 72.77±0.93% (wt.) was obtained with pyrolysis reactor at 450 °C, catalyst bed at 350 °C, and catalyst-to-tallow kernel oil ratio of 1:2. The GC-MS analysis of bio-oils identified the presence of alkanes, olefins, aromatics, esters, carboxylic acids, ketones, and alcohols.

Koul *et. al.* [13], investigated the effect of catalyst-to-biomass ratio on the yield of bio-oil organic phase by the catalytic upgrading of vapors from castor seed oil pyrolysis over different catalysts (Kaolin, CaO, ZnO), using fixed bed reactor in a two-stage reactor schema in laboratory scale. In addition, for the experiments giving the highest yields, the physicochemical properties including pH and kinematic viscosity were determined. The fixed bed pyrolysis reactor ($\phi_{id} = 6$ cm, $H = 21$ cm, $V_{Reactor} = 593.7$ mL), containing the castor seed oil (40 g), is separated 2.0 cm from the catalyst bed using glass wool in between. The experiments were carried out at 550 °C, heating rate of 25 °C/min, using 5, 10, 15, and 20 % (wt.) catalyst. The results show that the highest bio-oil yields were obtained using 15% (wt.) Kaolin, 15% (wt.) CaO, and 10% (wt.) ZnO. The highest bio-oil yields were 64.9, 66.4, and 65.8% (wt.) using Kaolin, CaO and ZnO, respectively, showing pH values of 8.36, 9.25 and 8.32, while the measured kinematic viscosity were 39.0, 8.3 and 28.0 (mm²/s). The FT-IR analysis identified the presence of hydrocarbons (alkanes, alkenes, and alkynes), as well as oxygenates (acids, aldehydes, ketones, esters, and amines). The GC-MS analysis of bio-oils obtained by catalytic upgrading of castor seed pyrolytic vapors using 15% (wt.) Kaolin, 15% (wt.) CaO, and 10% (wt.) ZnO identified the presence of aromatics, esters, amines, acids, alkanes, amides, alkenes, alcohols, ethers, as well as non-identified compounds, showing hydrocarbons contents (aromatics, alkanes, alkenes) of 21.92 (area.%), 12.43 (area.%), and 37.49 (area.%), respectively.

The work aims to investigate systematically the effect of catalyst-to-Tallow kernel oil ratio (0.0, 0.05, 0.075, 0.100) and reaction time (50, 60, 70, 80, 90, 100 and 120 minutes) by catalytic upgrading of pyrolysis vapors of residual fat at 400 °C and 1.0 atmosphere, on the yield, physicochemical properties (density, kinematic viscosity, and acid value) and chemical composition of hydrocarbons and oxygenates, of bio-oils, over a catalyst fixed bed reactor containing 0.0, 5.0, 7.5, and 10.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, in semi-pilot scale, using a process schema consisting of a two-stage reactor, whereas the first stage is a pyrolysis reactor of 2.0 L coupled to the second-stage, a catalyst fixed bed reactor of 53 mL, in batch mode.

2. Materials and Methods

2.1. Methodology

The process flowsheet illustrated in Figure 1 summarizes the applied methodology, described in a logical sequence of ideas, methods, and procedures to produce liquid mixtures rich in hydrocarbons by catalytic upgrading of residual fat pyrolysis vapors/volatiles at 400 °C and 1.0 atm, over a catalyst fixed bed reactor, using activated carbon pellets impregnated with sodium hydroxide, in semi-pilot scale. Initially, the residual fat is collected. Afterwards, it is subjected to pre-treatment (sieving) and separation processes (evaporation, dehydration). The activated carbon pellets impregnated with NaOH. Afterwards, the impregnated carbon pellets were washed with deionized water and submitted to drying followed by thermal activation. The experiments were carried out using a two-stage process schema consisting of a pyrolysis reactor and a catalyst fixed bed reactor. The effect of catalyst-to-residual fat and reaction time were analyzed. The physical-chemistry (density, kinematic viscosity, and acidity) properties and composition of hydrocarbons and oxygenates determined.

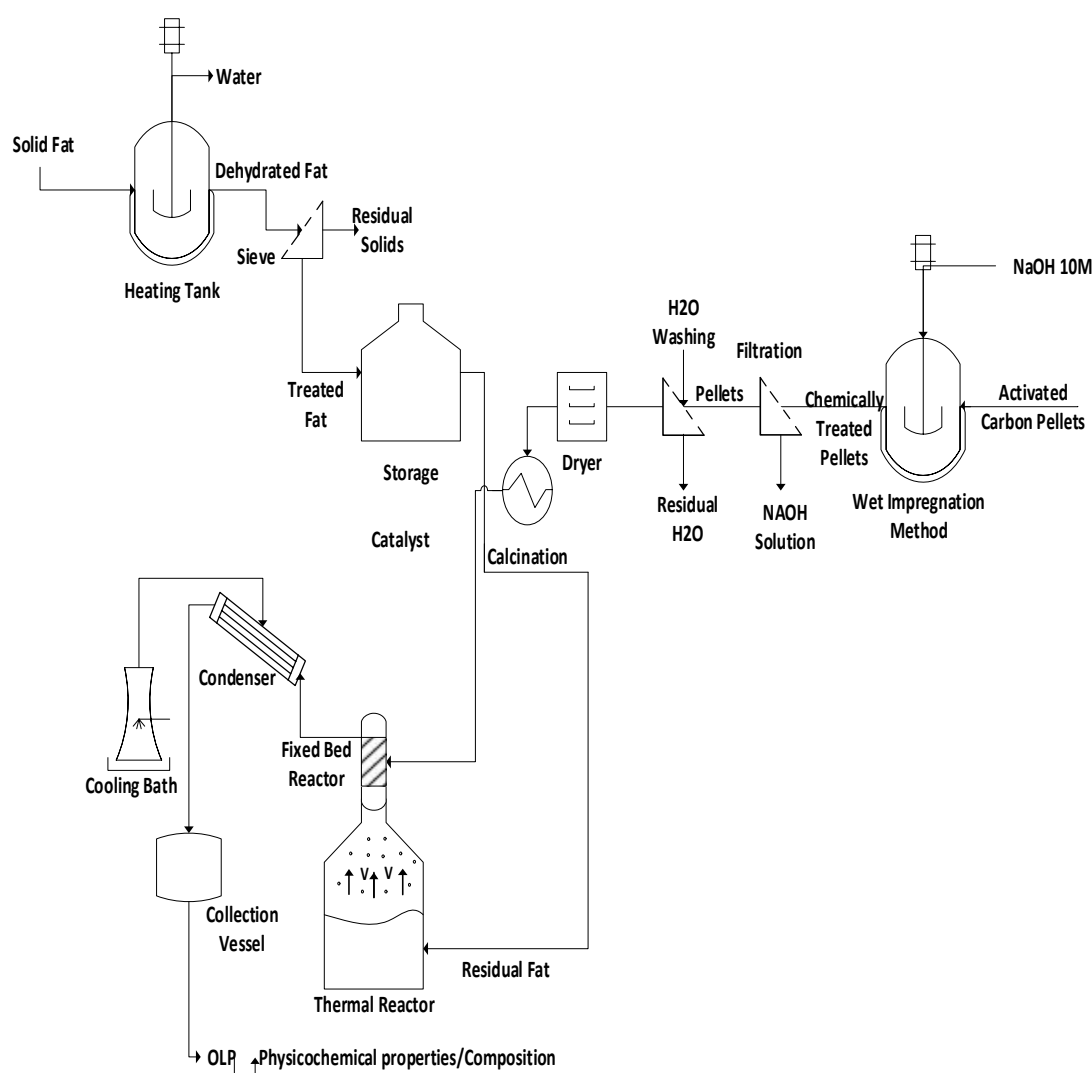


Figure 1. Process flowsheet by catalytic upgrading of residual fat vapors at 400 °C, 1.0 atm, 0.0, 5.0, 7.5, and 10.0% (wt.) of activated carbon pellets impregnated with 10.0 M NaOH, using a catalyst fixed bed reactor, in semi-pilot scale.

2.2. Materials

The residual fat was collected from a system of fat-retaining boxes at University Restaurant of UFPA. The residual fat was submitted to pre-processing (heating, homogenization) and separation processes (dehydration, sieving), as described elsewhere [14-16]. By the pre-processing, the fat was subjected to heating and mechanical agitation in order to evaporate the excess moisture and homogenize the lipid-base mixture. Afterwards, the liquified fat material was sieved to remove suspended and undesired solid materials, as described synthetically in flowsheet of Figure 1. The residual fat after pre-processing is illustrated in Figure 2.



Figure 2. Pre-treated residual fat used as feed material by catalytic upgrading of residual fat vapors at 400 °C, 1.0 atm, 0.0, 5.0, 7.5, and 10.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, using a catalyst fixed bed reactor, in semi pilot scale.

2.3. Chemical activation of commercial activated carbon pellets

The pellets of activated carbon (General Carbon Corp., New Jersey, GC C-40), chemically activated/impregnated with 10.0 M NaOH, were used as catalysts. The specifications of commercial activated carbon pellets are described in Table 1.

Table 1. Specifications of activated carbon GC C-40.

Specifications GC C-40	Units
Particle size (mm)	3.9-4.1
Mean particle diameter (mm)	4.0
CCl ₄ activity (%)	60 (minimum)
Butane activity (%)	23 (typical)
Iodine number (mg/g)	900
Surface area (m ² /g)	900
Hardness (%)	95
Moisture (%)	5.0 (maximum)
Density (g/cm ³)	0.45-0.55
pH	9.0-11.0

2.3.1. Wet impregnation, drying, and calcination

The activated carbon pellets were chemically impregnated by the wet impregnation method using a 10.0 M NaOH solution. The chemical impregnation of activated carbon pellets takes place in several steps, as illustrated in Figure 1. In this sense, a more detailed description of each pre-treatment step (wet impregnation, drying, and calcination) should be highlighted for a better understanding as follows. First, 500 g of activated carbon pellets were placed in contact

with 320 g of a 10.0 M NaOH solution for 4.0 hours at 30 °C. Afterwards, the solution was drained using a filter. Then, the moist impregnated pellets washed with deionized water, to remove the excess NaOH, and the procedure repeated until stable pH value of 10.0 was achieved. Afterwards, the pellets were subjected to drying using an oven (Model Bio SEA - 40L) for a period of 12 hours, at a temperature of 105 °C, to remove the excess water. Finally, the activated carbon pellets impregnated with NaOH were calcined at 600 °C for 03 hours using a muffle (Model Bio SEA - 40L). A total of 04 batches were carried out. The commercial activated carbon before and after impregnation with sodium hydroxide is shown in Figure 3.

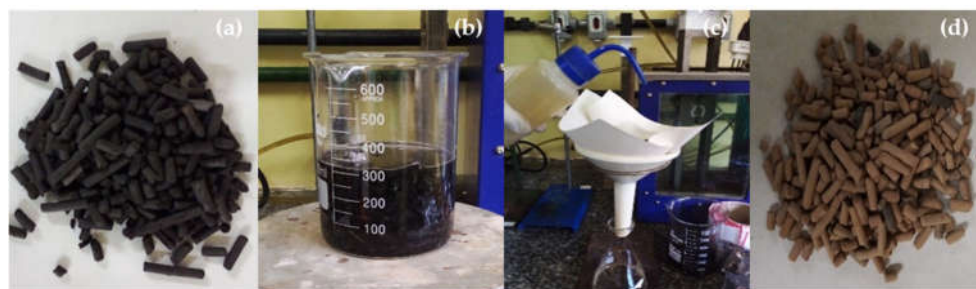


Figure 3. Activated carbon pellets impregnated with a 10.0 M NaOH solution [Activated Carbon Pellets (a); Wet Impregnation (b); Water Washing (c); Activated Carbon Pellets Impregnated with NaOH (d)].

2.4. Characterization of residual fat

The pre-treated residual fat was physicochemical characterized for density, kinematic viscosity, refractive index and acid value according to official methods AOCS Cc 10c-95, ASTM D 2515, AOCS Cc 7-25, and AOCS Cd3d-63 [14-18].

2.5. Experimental apparatus and procedures

2.5.1. Experimental apparatus

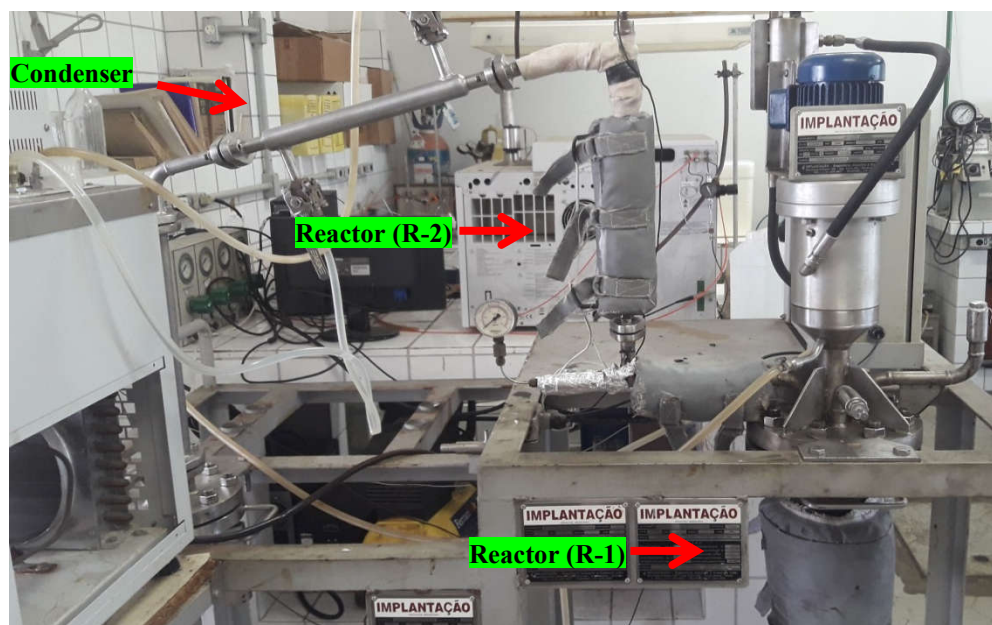


Figure 4. Thermal catalytic cracking unit with a pyrolysis reactor (R-1) and a fixed bed reactor (R-2) in bench scale.

The thermal catalytic cracking unit with a fixed bed reactor in semi pilot, depicted in Figure 4, described in details elsewhere [15]. By the thermal catalytic

cracking experiments, at the output of pyrolysis/batch reactor (R-1), a second AISI 304 stainless steel fixed bed reactor (R-2) is coupled. The fixed bed reactor (R-2) of cylindrical geometry has 30 cm height and 15 mm internal diameter ($V_{(R-2)} = 53$ ml). An spiral-shaped electrical resistance, with 1.5 kW, was inserted around the reactor (R-2). A glass wool thermal blanket was used as a thermal insulator in order to avoid energy loss to the environment. A type K, sheath thermocouple (550 °C) is placed inside the reactor (R-2) for measuring its temperature. Figure 5 shows the schematic diagram of bench scale catalytic cracking unit with a fixed bed (R-2).

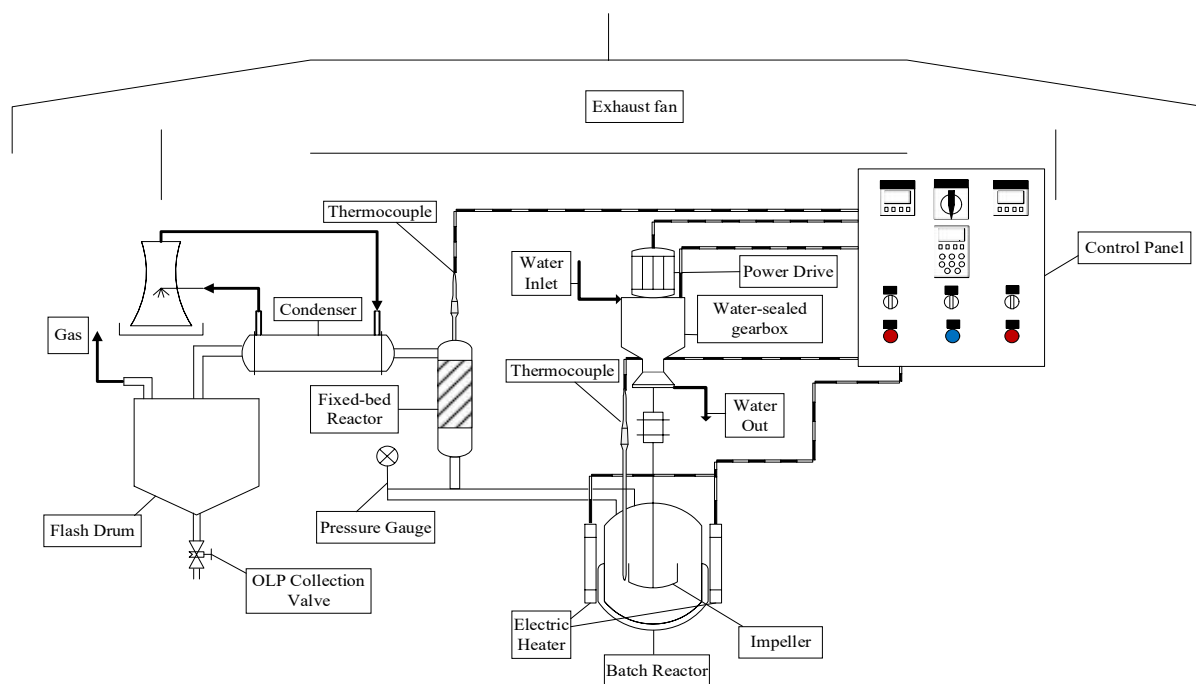


Figure 5. Schematic diagram of bench scale stainless steel catalytic cracking reactor with a pyrolysis/batch reactor and a fixed bed reactor.

2.5.2. Experimental procedures

• 2.5.2.1 Pyrolysis

By the thermal cracking (pyrolysis) of residual fat, the residual fat was weighed (1000 g) using an electronic balance (Mars, AL500). Afterwards, the residual fat is added within the stirred tank reactor (R-1), whereas pyrolysis takes place. After sealing the reactor, the experimental apparatus has been set up. Then, the cooling system is turned on and the water temperature was set at 10 °C. Afterwards, the programming step is carried out, thus making it possible to set the desired operating parameters, including heating rate (10 °C/min), cracking temperature (400 °C) and the mechanical impeller speed (100 rpm). After the experiment starts, the operational parameters (heating rate, cracking temperature, and mechanical impeller speed) were recorded every 10 minutes. By each thermal cracking experiment, 03-04 (three/four) samples were withdrawn along with the reaction time in order to study and/or investigate the reaction kinetics of organic liquid products. In addition, the non-condensable gases were burned at the exit of the gas pipeline line (gas flare). The mass of solid phase products (coke) was collected and weighed, and the mass of gas computed by difference. The liquid phase products were submitted to pre-treatments of decantation to remove water and filtration. Then, organic liquid products characterized for density, kinematic viscosity, refractive index, and acid value.

• 2.5.2.2 Thermal catalytic cracking

By the thermal catalytic cracking experiments, the activated carbon pellets impregnated with 10.0 M NaOH were placed inside the fixed bed reactor (R-2). Afterwards, the reactor (R-2) was vertically coupled between the reactor (R-1) and the condenser, as shown in Figure 4. Then, the desired operating parameters are set up, including the heating rate (10 °C/min), cracking temperatures (400 °C) of both reactors (R-1) and (R-2), as well as the mechanical impeller speed (100 rpm). Furthermore, the temperature of fixed bed reactor (R-2) was set up approximately 10 °C/20 °C higher than that of reactor (R-1) in order to avoid condensation of ascending vapors produced inside the reactor (R-1) through the porous of pellets catalyst and between the void volume of fixed bed reactor (R-2). That is, thermal cracking of residual fat takes place in reactor (R-1), producing an ascending vapor phase (gas phase reaction products) that flows through the porous of pellets catalyst and between the void volume of fixed bed reactor, whereas a heterogeneous gas-solid reaction takes place in catalyst fixed bed (R-2). The thermal catalytic cracking experiments carried out with 5.0, 7.5, and 10.0% (wt.) chemically activated carbon pellets inside the fixed bed reactor (R-2). By each thermal catalytic cracking experiment, 03-04 (three/four) samples were withdrawn along with the reaction time in order to study and/or investigate the reaction kinetics of organic liquid products. In addition, the non-condensable gases were burned at the exit of the gas pipeline line (gas flare). The mass of solid phase products (coke) was collected and weighed, and the mass of gas computed by difference. The liquid phase products were submitted to pre-treatments of decantation to remove water and filtration. Then, organic liquid products physicochemical characterized for density, kinematic viscosity, refractive index, and acid value.

2.6. Physicochemical and chemical composition of bio-oil

2.6.1. Physicochemical characterization of bio-oil and aqueous phase

The bio-oil was characterized for acidity, density at 25 °C, kinematic viscosity, and refractive index according to AOCS and ASTM methods, as described in the literature [14-18].

2.6.2. Chemical composition of bio-oil and aqueous phase

The chemical composition of bio-oil and aqueous phase were determined by GC-MS and the equipment and procedure described in details by Castro *et al.* [17]. The concentrations were expressed in area, as no internal standard was injected for comparison the peak areas.

2.7. Characterization of activated red mud pellets

2.7.1. SEM and EDX analysis

The morphological characterization of activated carbon pellets impregnated with 10.0 M NaOH was performed by scanning electron microscopy using a microscope (Tescan GmbH, Czech Republic, Model: Vega 3). The samples were covered with a thin layer of gold using a Sputter Coater (Leica Biosystems, Germany, Model: Balzers SCD 050). Elemental analysis and mapping were carried out by energy dispersive x-ray spectroscopy (Oxford instruments, UK, Model: Aztec 4.3).

2.7.2. XRD analysis

The crystalline characterization of activated carbon pellets impregnated with 10.0 M NaOH was performed by x-ray diffraction using a diffractometer (Rigaku, Japan, Model: MiniFlex600) at the Laboratory of Structural Characterization (FEMAT/UNIFESSPA) and the equipment specifications described as

follows: *generator* (maximum power: 600 W; tube voltage: 40 kV; tube current: 15 mA; X-ray tube: Cu), *optics* (fixed divergence, scattering and receiving slit; filter; K β sheet; monochromator: graphite; soller slit: 5.0°), *goniometer* (model: vertical, radius: 150 mm, scanning range: -3 A, 145° (2 θ); scanning speed: 0.01 to 100°/min (2 θ); accuracy: $\pm 0.02^\circ$) and *detector* (high-speed silicone tape).

2.8. Mass balances by catalytic cracking of vapor phase products

Application of mass conservation principle in the form an overall mass balance within the pyrolysis/catalytic reactor, operating in batch mode, open thermodynamic system, yields the following equations for the material system R-1 (pyrolysis reactor).

$$m_{in,pyrolysis} - m_{out,pyrolysis} = \frac{dm_{Feed}}{d\tau} \quad (1)$$

$$m_{in,pyrolysis} = 0 \quad (2)$$

$$-m_{out,pyrolysis} = \frac{dm_{Feed}}{d\tau} \quad (3)$$

$$-m_{out,pyrolysis} = m_{vapors,pyrolysis} \quad (4)$$

Where m_{in} is the mass flow rate entering reactor R-1, $m_{out,pyrolysis}$ is the mass flow rate leaving the reactor R-1, $\frac{dm_{Feed}}{d\tau}$ is the time rate variation of feed mass inside reactor R-1, $m_{vapors,pyrolysis}$ is the mass flow rate of pyrolysis vapors/volatiles leaving the reactor R-1 and entering the reactor R-2 (catalyst bed reactor). By applying an overall steady state mass balance within the catalyst reactor, operating in continuous mode, open thermodynamic system, yields the following equations for the material system R-2 (catalyst reactor). Assuming that all the vapors leaving reactor R-1, reacts at the solid pellets surface producing gaseous species, the following equations for the material system R-2 (catalyst reactor) applies.

$$m_{vapors,pyrolysis} - m_{out,catalyst} = \frac{dm_{vapors}}{d\tau} \quad (5)$$

$$\frac{dm_{vapors}}{d\tau} = 0 \quad (6)$$

$$m_{vapors,pyrolysis} - m_{out,catalyst} = 0 \quad (7)$$

$$m_{out,catalyst} = m_{gas} + m_{bio-oil} \quad (8)$$

Where $m_{vapors,pyrolysis}$ is the mass flow rate entering reactor R-2, $m_{out,catalyst}$ is the mass flow rate leaving the reactor R-2, m_{gas} is the mass flow rate of non-condensable gases leaving the reactor R-2, computed by difference, and $m_{bio-oil}$ is the mass flow rate of bio-oil leaving the condenser. The mass of solid remaining in the reactor R-1 is m_{solid} . By performing a steady state global mass balance within the control volume consisting of reactors R-1 and R-2 yields equation (9).

$$m_{Feed} = m_{solid} + m_{gas} + m_{bio-oil} \quad (9)$$

The process performance evaluated by computing the yields of bio-oil, solid (coke), and gas defined by equations (10) and (11), and the yield of gas by difference, using equation (12).

$$Y_{bio-oil}[\%] = \frac{M_{bio-oil}}{M_{Feed}} \times 100 \quad (10)$$

$$Y_{solids}[\%] = \frac{M_{solids}}{M_{Feed}} \times 100 \quad (11)$$

$$Y_{gas}[\%] = 100 - (Y_{bio-oil} + Y_{solids}) \quad (12)$$

4. Results

4.1. Characterization of catalyst

4.1.1. SEM analysis

The microscopies of activated carbon pellets (General Carbon Corp., New Jersey, GC C-40) depicted in Figure 6 (c) show its highly porous structure, being the SEM images according to those reported by Thakur *et. al.* [19]. After calcination of activated carbon pellets impregnated with 10.0 M NaOH, SEM images show the appearance of cavities with sizes between 4.0 and 10.0 μm as shown in Figure 7 (c), proving that high concentrated NaOH solutions has caused a great change on the activated carbon pellets morphology and texture. The SEM images of activated carbon pellets impregnated with 10.0 M NaOH after upgrading of residual fat pyrolysis vapors at 400 $^{\circ}\text{C}$, 1.0 atmosphere, using a semi-pilot scale two-stage reactor of 2.0 L, illustrate that porous and cavities were filled with molecules of carbon due to gas-solid reactions on the surface and porous structure of activated carbon pellets, as shown in Figure 8 (c), corroborated in Table 2, as coke deposits may be formed on the surface of catalysts by polymerization of aromatics present in the pyrolysis vapors or formed during reaction on the gas-solid interface [20]. In fact, activated carbon pellets impregnated with 10.0 M NaOH after upgrade of residual fat pyrolysis vapors became black due to the carbonization that takes place inside the pores within the solid surface. According to Santillan-Jimenez and Crocker [21], by the catalytic deoxygenation of fatty acids into hydrocarbons via decarboxylation and/or decarbonylation, CO_2 is formed by decarboxylation, while H_2O and CO are formed by decarbonylation, being those compounds released into the gaseous phase.

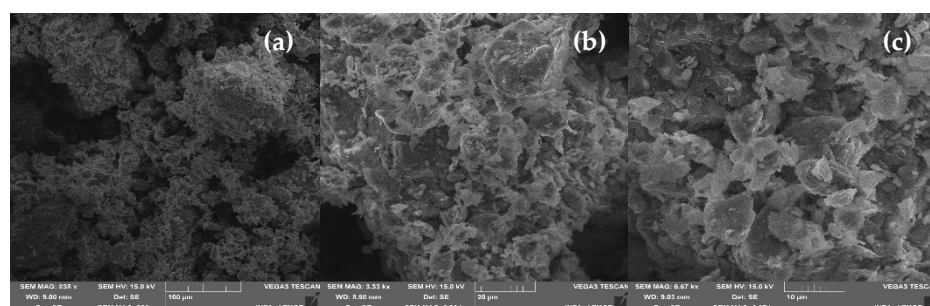


Figure 6. SEM of activated carbon pellets [MAG: 838 x (a); MAG: 3.33 kx (b); MAG: 6.67 kx (c)].

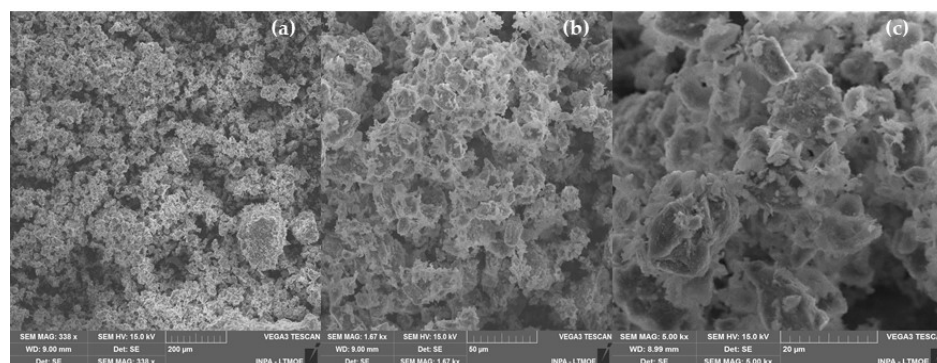


Figure 7. SEM of activated carbon pellets impregnated with 10.0 M NaOH [MAG: 338 x (a); MAG: 1.67 kx (b); MAG: 5.00 kx (c)].

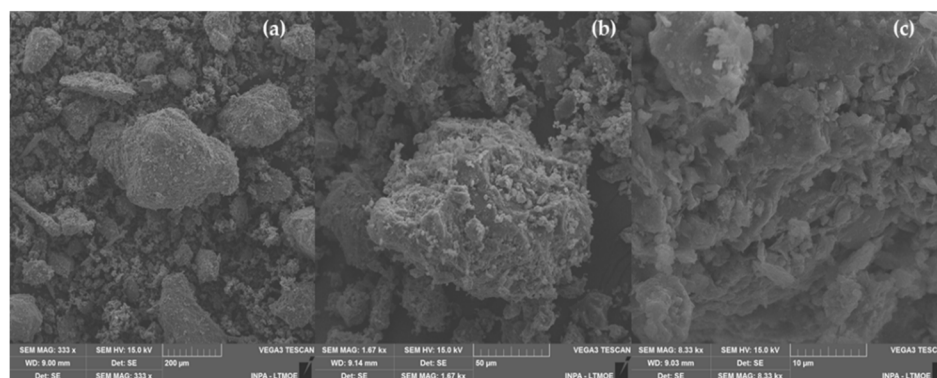


Figure 8. SEM of activated carbon pellets after catalytic upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atm, with 5.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, in batch mode, using a semi-pilot scale reactor of 2.0 L [MAG: 838 x (a); MAG: 3.33 kx (b); MAG: 6.67 kx (c)].

4.1.2. EDX analysis

Table 2 shows the results of elemental analysis performed by energy dispersive x-ray spectroscopy at a point for activated carbon pellets, activated carbon pellets impregnated with 10.0 M NaOH after calcination at 600 °C, and activated carbon pellets impregnated with 10.0 M NaOH after upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atmosphere, using a semi-pilot scale two-stage reactor of 2.0 L. The impregnation of activated carbon pellets with sodium hydroxide had little effect on the elemental analysis, except on the elemental analysis of *Silicium* (Si) and *Sodium* (Na). The mass (wt.%) of *Silicium* (Si) decreases after impregnation with a 10.0 M NaOH, being according to the results reported by Fertani-Gmati *et. al.* [22], as silicates dissolves in NaOH [$\text{SiO}_2 + 6(\text{NaOH})_{\text{H}_2\text{O(L)}} \rightarrow (\text{Si}_4\text{O}_{11}\text{Na}_6)_{\text{H}_2\text{O(L)}} + 3\text{H}_2\text{O}$]. In addition, the impregnation of activated carbon pellets with NaOH caused an increase on the mass (wt.%) of *Sodium* (Na), as shown in Table 2. After upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atmosphere, over a catalyst fixed bed reactor packed with activated carbon chemically activated with 10.0 M NaOH, using a semi-pilot scale two-stage reactor of 2.0 L, one observes an increase on the mass (wt.%) of carbon probably due not only to the adsorption of carbon molecules within the pores of catalyst surface, but also to formation of coke deposits on the surface of catalysts by polymerization of aromatics present in the pyrolysis vapors or formed during reaction on the gas-solid interface [20], as gas-solid reactions takes place.

Table 2. Percentages in mass and atomic mass of activated carbon pellets, activated carbon pellets impregnated with 10.0 M NaOH after calcination at 600 °C and 03 hours, and activated carbon pellets impregnated with 10.0 M NaOH after upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atmosphere, using a semi-pilot scale two-stage reactor of 2.0 L.

Chemical Elements	Catalyst								
	Activated carbon pellets			Activated carbon pellets impregnated with 10.0 M NaOH			Activated carbon pellets impregnated with 10.0 M NaOH after 400 °C		
	Mass [wt.%]	Atomic Mass [wt.%]	SD	Mass [wt.%]	Atomic Mass [wt.%]	SD	Mass [wt.%]	Atomic Mass [wt.%]	SD
C	-	-	-	-	-	-	43.28	56.58	0.37
O	58.21	72.00	0.11	61.67	74.58	0.08	31.67	31.08	0.30
Mg	0.94	0.77	0.02	1.01	0.80	0.02	0.30	0.19	0.02
Al	10.93	8.02	0.04	10.77	7.73	0.03	4.20	2.44	0.04
Si	22.49	15.85	0.07	17.88	12.32	0.04	8.67	4.85	0.07
K	2.05	1.04	0.02	1.93	0.96	0.01	1.18	0.47	0.02
Ca	1.40	0.69	0.01	1.72	0.83	0.01	4.84	1.90	0.04
Fe	3.07	1.09	0.02	2.20	0.76	0.01	3.50	0.98	0.04
Ti	0.45	0.19	0.01	0.49	0.20	0.01	0.23	0.08	0.01
Na	0.40	0.34	0.03	1.82	1.53	0.02	2.05	1.40	0.04
Mn	0.06	0.02	0.01	0.05	0.02	0.01	0.08	0.02	0.01

SD= Standard Deviation

4.1.3. XRD analysis

The XRD of activated carbon pellets impregnated with 10.0 M NaOH after calcination at 600 °C, 03 hours, is shown in Figure 9. The diffractogram shows the presence of 03 (three) crystalline phases, Quartz (SiO_2), Calcite (CaCO_3), and Sodium Oxide Hydrate ($\text{H}_2\text{Na}_2\text{O}_2$). For the crystalline phase Quartz, the XRD identified the presence of 02 (two) peaks, one of high intensity, observed on the position 2θ : 26.6 (100%) and another of low intensity on the position 2θ : 20.8 (21.6%). The phase Calcite (CaCO_3) of Rhombohedral structure was observed in 02 (two) peaks, one of high intensity, observed on the position 2θ : 29.4 (100%) and another of low intensity on the position 2θ : 48.6 (19.1%). The phase Sodium Oxide Hydrate has been observed in 02 (two) peaks of intensity on the position 2θ : 37.1 and 2θ : 40.8, showing that impregnation of activated carbon pellets with NaOH was effective.

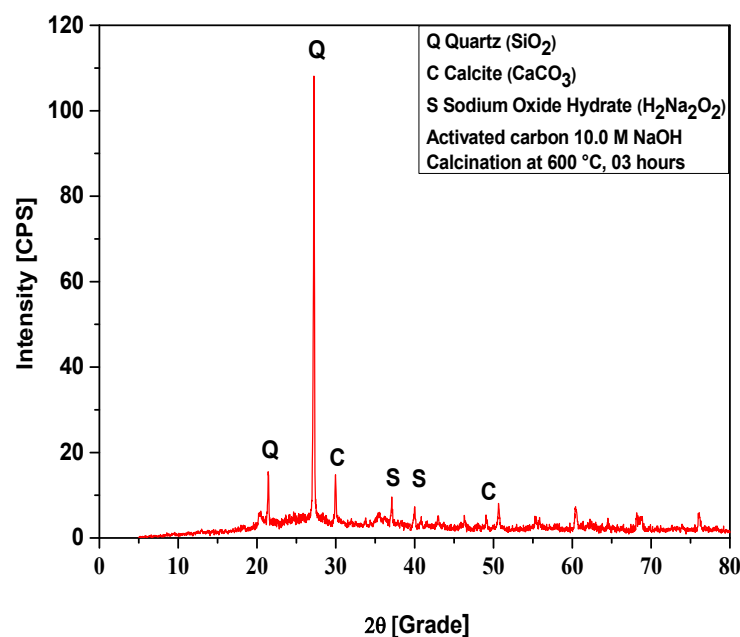


Figure 9. XRD of activated carbon pellets impregnated with 10.0 M NaOH after calcination at 600 °C, 03 hours.

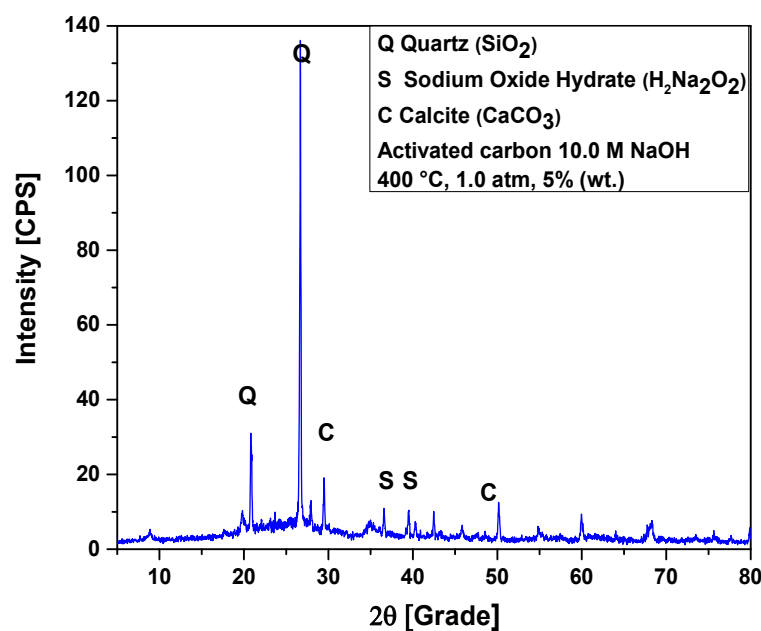


Figure 10. XRD of activated carbon pellets impregnated with 10.0 M NaOH after the upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atmosphere, 5.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, using a catalyst fixed bed reactor, in a semi-pilot scale two-stage reactor of 2.0 L.

The XRD of activated carbon pellets impregnated with 10.0 M NaOH after the upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atmosphere, 5.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, using a catalyst fixed bed reactor, in a semi-pilot scale two-stage reactor of 2.0 L, is shown in

Figure 10. The diffractogram shows the presence of 03 (three) crystalline phases, Quartz (SiO_2), Calcite (CaCO_3), and Sodium Oxide Hydrate ($\text{H}_2\text{Na}_2\text{O}_2$). The XRD identified the presence of 02 (two) peaks, one of high intensity, observed on the position 2θ : 26.6 (100%) and another of low intensity on the position 2θ : 20.8 (21.6%), associated to the crystalline phase Quartz (SiO_2). The peaks associated to Sodium Oxide Hydrate ($\text{H}_2\text{Na}_2\text{O}_2$) were observed on the position 2θ : 37.1 and 2θ : 40.8. The phase Calcite (CaCO_3) of Rhombohedral structure was observed in 02 (two) peaks, one of high intensity, observed on the position 2θ : 29.4 (100%) and another of low intensity on the position 2θ : 48.6 (19.1%).

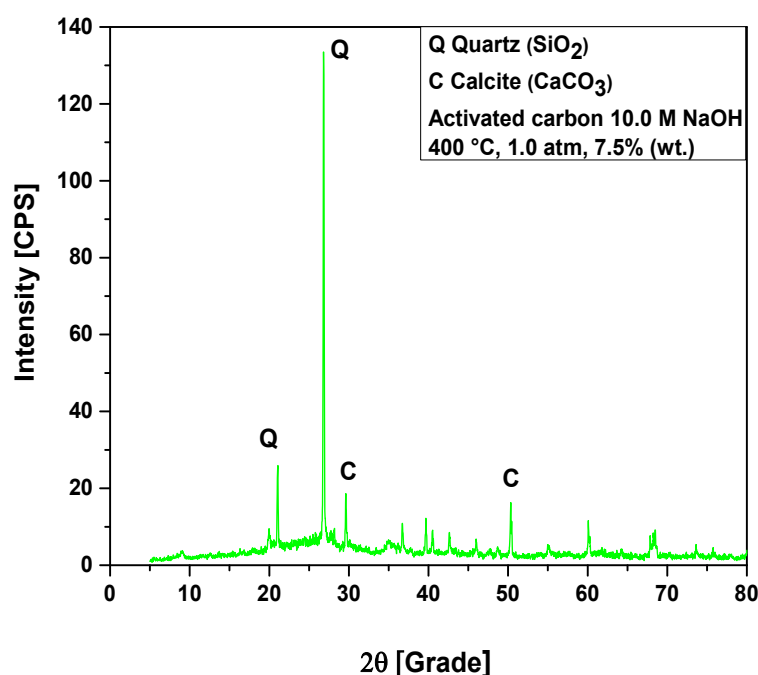


Figure 11. XRD of activated carbon pellets impregnated with 10.0 M NaOH after the upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atmosphere, 7.5% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, using a catalyst fixed bed reactor, in a semi-pilot scale two-stage reactor of 2.0 L.

The XRD of activated carbon pellets impregnated with 10.0 M NaOH after the upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atmosphere, 7.5% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, using a catalyst fixed bed reactor, in a semi-pilot scale two-stage reactor of 2.0 L, is shown in Figure 11. The diffractogram shows the presence of 02 (two) crystalline phases, Quartz (SiO_2) and Calcite (CaCO_3). The XRD identified the presence of 02 (two) peaks, one of high intensity, observed on the position 2θ : 26.7 (100%) and another of low intensity on the position 2θ : 20.9 (21.4%), associated to the crystalline hexagonal phase Quartz (SiO_2). The phase Calcite (CaCO_3) of Rhombohedral structure was observed in 02 (two) peaks, one of high intensity, observed on the position 2θ : 29.6 (100%) and another of low intensity on the position 2θ : 48.7 (15.0%).

The XRD of activated carbon pellets impregnated with 10.0 M NaOH after the upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atmosphere, 10.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, using a catalyst fixed bed reactor, in a semi-pilot scale two-stage reactor of 2.0 L, is shown in Figure 12. The diffractogram shows the presence of 02 (two) crystalline phases, Quartz (SiO_2) and Calcite (CaCO_3). The XRD identified the presence of 02 (two)

peaks, one of high intensity, observed on the position 2θ : 26.6 (100%) and another of low intensity on the position 2θ : 20.8 (21.6%), associated to the crystalline hexagonal phase Quartz (SiO_2). The phase Calcite (CaCO_3) of Rhombohedral structure was observed in 02 (two) peaks, one of high intensity, observed on the position 2θ : 29.4 (100%) and another of low intensity on the position 2θ : 48.6 (19.1%).

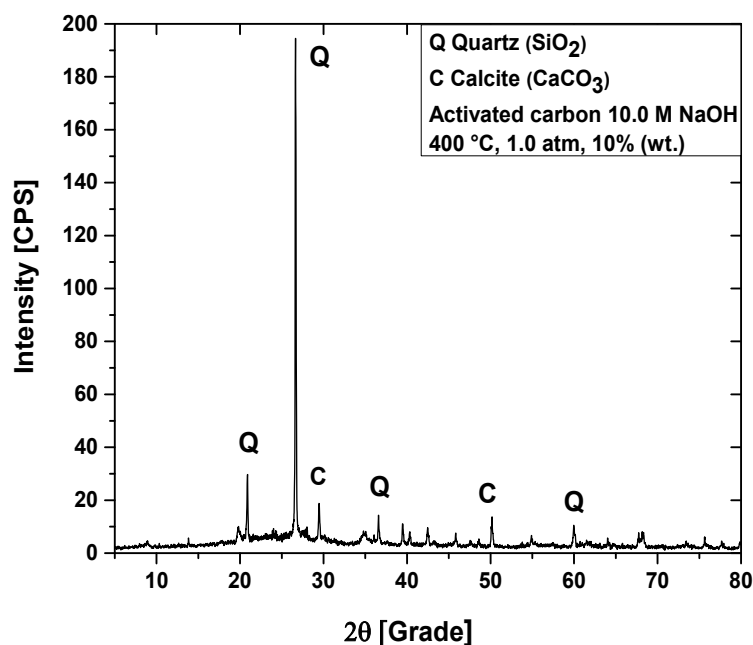


Figure 12. XRD of activated carbon pellets impregnated with 10.0 M NaOH after the upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atmosphere, 10.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, using a catalyst fixed bed reactor, in a semi-pilot scale two-stage reactor of 2.0 L.

4.2. Upgrading of residual fat pyrolysis vapors over activated carbon pellets

4.2.1. Process conditions, mass balances, and yields of reaction products

Table 3. Process parameters, mass balances, and yields of reaction products (liquids, solids, H₂O, and gas) by catalytic upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atm, 0.0, 5.0, 7.5, and 10.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, using a catalyst fixed bed reactor, in semi pilot scale.

Process Parameters	400 (°C)			
	0.0 (wt.)	5.0 (wt.)	7.5 (wt.)	10.0 (wt.)
Mass of residual fat (g)	1200	700.4	700	700
Cracking time (min)	90	100	90	120
Initial cracking temperature (°C)	360	370	395	396
Mechanical system stirring speed (rpm)	100	600	600	600
Mass of solid (Coke) (kg)	50	95.0	68.25	73.8
Mass of liquid (Bio-oil) (kg)	917.0	389.11	353.07	211.55
Mass of H ₂ O (kg)	78.66	19.8	22.35	21.00
Mass of gas (kg)	154.34	196.49	256.36	393.65
Yield of Bio-oil (wt.%)	76.41	55.55	50.44	30.22
Yield of H ₂ O (wt.%)	6.55	2.83	3.19	3.00
Yield of Coke (wt.%)	4.17	13.56	9.75	11.18
Yield of Gas (wt.%)	12.87	27.89	36.62	55.60

Table 3 shows bio-oil yields between 30.22 and 76.41% (wt.), aqueous phase yields between 2.83 and 6.55% (wt.), solid phase yields between 4.17 and 13.56% (wt.), and gas yields between 12.87 and 55.60% (wt.). The yields of bio-oil are lower than those reported by Yu *et al.* [12], for the catalytic upgrade of Tallow kernel oil over SiC foam-MCM41 and by Koul *et al.* [13], for the catalytic upgrading of vapors from *castor seed oil* pyrolysis over Kaolin, CaO, and ZnO, but higher than those reported by Wang *et al.* [2], for catalytic upgrading of vapors from *Chromolaena odorata* and *soybean soapstock*, by Vichaphund *et al.* [6], for the catalytic upgrading of *Jatropha waste* over HZSM-5, by Wang *et al.* [7], for catalytic upgrading of *soapstock* vapors over HZSM-5, and by Wu *et al.* [8], for the catalytic upgrading of *soapstock* and *straw* vapors using HZSM-5 and MCM-4. The yields of bio-oil by the catalytic upgrading of lipid-base materials stays around 30.0 and 60.0% (wt.) [2, 6-8, 12-13].

4.2.2. Effect of catalyst-to-residual fat ratio on the yield of bio-oil

Figures 13 shows the effect of catalyst-to-biomass ration on the yields of bio-oil by the upgrading of residual fat pyrolysis vapors/volatiles at 450 °C, 1.0 atm, 0.0, 5.0, 7.5, and 10.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, using a catalyst fixed bed reactor, in a semi-pilot scale two-stage reactor of 2.0 L, compared with similar studies in the literature [2, 6-8, 12-13].

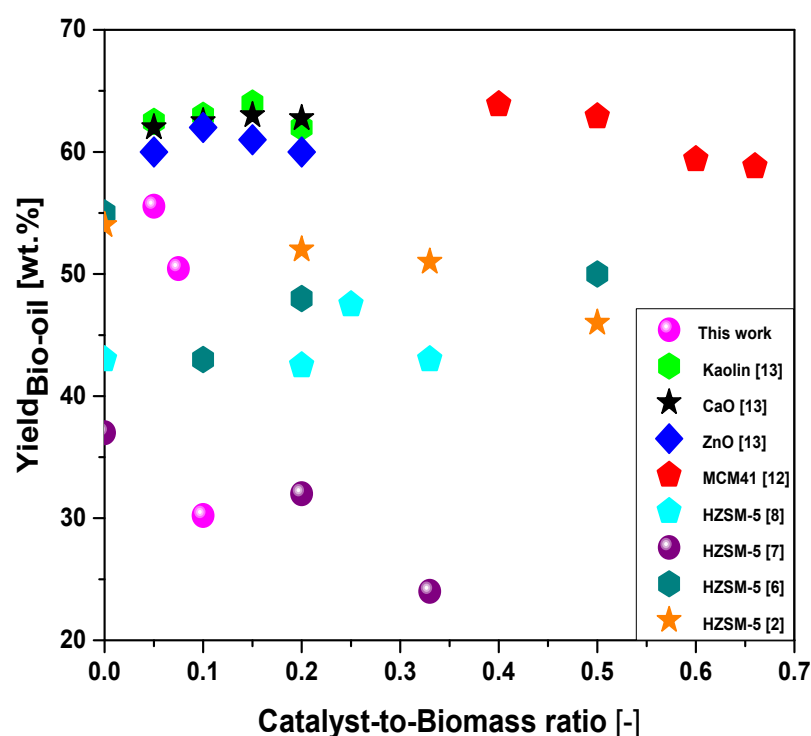


Figure 13. Effect of catalyst-to-biomass ratio within the range 0.0-0.7 on the yield of bio-oil by the upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atmosphere, 5.0, 7.5, and 10.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, using a catalyst fixed bed reactor, in a semi-pilot scale two-stage reactor of 2.0 L, compared with similar studies in the literature [2, 6-8, 12-13].

By increasing the catalyst-to-residual fat ratio causes a decrease on the yield of bio-oil. The results are according to similar studies reported by Koul *et al.* [13], for the catalytic upgrading of vapors from *castor seed oil* pyrolysis over ZnO, between catalyst-to-biomass ratio of 0.1 and 0.2, by Yu *et al.* [12], for the catalytic upgrade of Tallow kernel oil over SiC foam-MCM41, Vichaphund *et al.* [6], for the catalytic upgrading of *Jatropha waste* over HZSM-5, by Wang *et al.* [7], for catalytic upgrading of *soapstock* vapors over HZSM-5, and by Wang *et al.* [2], for catalytic upgrading of vapors from *Chromolaena odorata* and *soybean soapstock*. In all studies [2, 6-7, 12-13], the feed materials have a great content of lipids and most the catalysts used were HZSM-5, except the ZnO [13]. In fact, ZnO reacts with carboxylic acids present in oils and fats, forming its corresponding carboxylates, that is, forming zinc soaps [23]. Probably, the Sodium Oxide Hydrate ($\text{H}_2\text{Na}_2\text{O}_2$), identified in the diffractogram of Figure 9, reacts with carboxylic acids present in the residual fat, forming sodium soaps. In fact, as observed by Yu *et al.* [12], higher catalyst-to-biomass ratios promotes the formation of gas, causing an opposite effect on the yield of bio-oil [2, 6-7, 12-13]. According to Duan *et al.* [24], increasing the catalyst-to-biomass ration increases the residence time of pyrolysis vapors inside the catalyst bed, increasing the rate of catalytic cracking of vapors/volatiles, thus promoting the formation of gaseous compounds. On the contrary, the yields of gas decrease with increasing catalyst-to-biomass, which is according to the literature [12], due probably to secondary cracking of pyrolysis vapors within the catalyst bed [23-24]. Finally, it is to expect that type of catalyst (acid or basic), catalyst morphological and textural properties and selectivity effects the formation of liquid (bio-oil), solid (coke) and gaseous

products by catalyst upgrade of pyrolysis vapors over a bed of catalyst coke [2, 6-10, 12-13]. Furthermore, the yield of reaction products and chemical composition of bio-oil are strongly dependent on process variables including reaction temperature [1-2, 7-8, 12], catalyst-to-biomass [1-4, 6-8, 12-13], residence time [7], and raw material composition [1-13].

4.2.3. Effect of reaction time on the physicochemical properties of bio-oil

Table 4. Effect of reaction time on the density, kinematic viscosity, and acidity of bio-oil by catalytic upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atm, 0.0, 5.0, 7.5, and 10.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, using a catalyst fixed bed reactor, in a semi-pilot scale two-stage reactor of 2.0 L.

Temperature/Catalyst	t_{Reaction} [min]	Physical chemistry properties		
		ρ [g/cm ³]	I.A [mg KOH/g]	ν [mm ² /s]
400 °C	50	0.8520	135.26	8.10
	60	0.8372	54.24	5.70
	70	0.8299	7.12	3.60
	80	0.8220	2.37	3.03
400 °C, 5% ACP (wt.)	50	0.9266	144.14	7.74
	60	0.8956	133.98	7.64
	80	0.8864	131.36	4.36
	100	0.8436	75.4	2.89
400 °C, 7.5% ACP (wt.)	70	0.8957	143.39	2.72
	90	0.8688	134.32	2.68
	120	0.8507	51.71	2.65
400 °C, 10% ACP (wt.)	60	0.8599	143.39	3.74
	70	0.8493	132.49	3.69
	90	0.8414	64.92	2.24

The effect of reaction time on the density, kinematic viscosity, and acidity of bio-oil by catalytic upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atm, 0.0, 5.0, 7.5, and 10.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, using a catalyst fixed bed reactor, in a semi pilot scale two-stage reactor of 2.0 L, is described in Table 4. The density, kinematic viscosity, and acidity of bio-oils decrease with increasing reaction time. This is because at the beginning of reaction, the primary cracking occurs, forming oxygenated compounds [14-16, 18]. Then, in the middle to the end of the reaction, secondary cracking occurs by de-carbonylation and de-carboxylation reactions, producing hydrocarbons [14-16, 18, 25-32], and hydrocarbons have superior physicochemical properties than oxygenates, that is, lowers density, lower kinematic viscosity and lower acidity.

• 4.2.3.1 Effect of reaction time on the density of bio-oil

The effect of reaction time on the density of bio-oil by catalytic upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atm, 0.0, 5.0, 7.5, and 10.0% (wt.) of activated carbon pellets impregnated with 10.0 M NaOH, using a catalyst fixed bed reactor, in semi pilot scale two-stage reactor of 2.0 L, is shown in Figure 14.

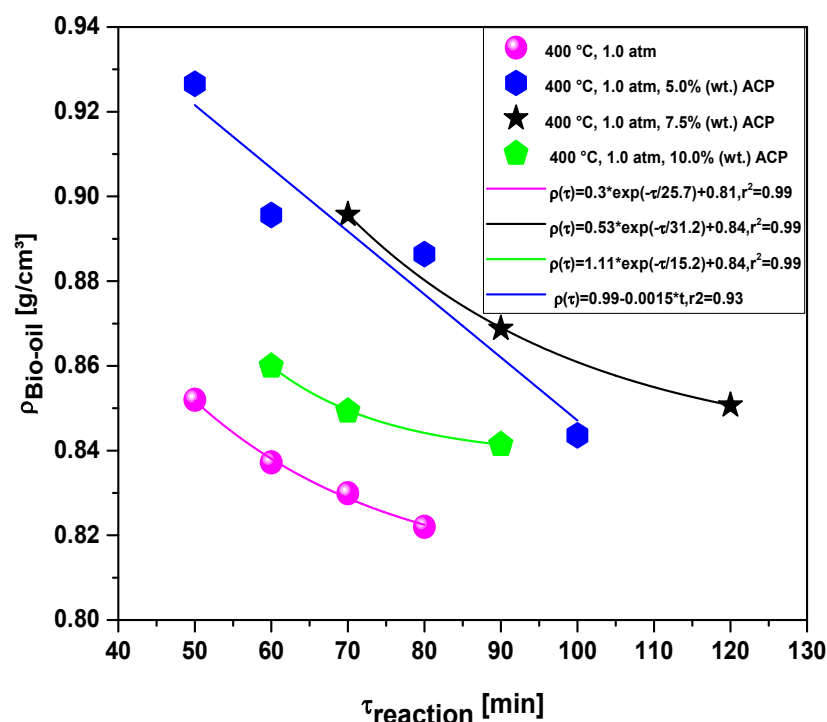


Figure 14. Effect of reaction time on the density of bio-oil by the upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atmosphere, 0.0, 5.0, 7.5, and 10.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, using a catalyst fixed bed reactor, in a semi-pilot scale two-stage reactor of 2.0 L.

For all the experiments, the density of bio-oil decreases with reaction time. For the experiments with 0.0, 7.5, and 10.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, the experimental data were correlated with a first order exponential decay model, exhibiting *root-mean-square error* (r^2) of 0.99, while the experiment with 5.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH was correlated with a 1st order linear equation, exhibiting *root-mean-square error* (r^2) of 0.93. The experiment without catalyst gives the lowest bio-oils densities, showing that a fixed bed filled with activated carbon pellets impregnated with 10.0 M NaOH promotes not only the formation of gas but also oxygenates compounds.

• 4.2.3.2 Effect of reaction time on the viscosity of bio-oil

Figure 14 illustrates the effect of reaction time on the kinematic viscosity of bio-oil by catalytic upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atm, 0.0, 5.0, 7.5, and 10.0% (wt.) of activated carbon pellets impregnated with 10.0 M NaOH, using a catalyst fixed bed reactor, in semi pilot scale two-stage reactor of 2.0 L. The kinematic viscosity of bio-oils decreases with increasing reaction time increases. The kinematic viscosity of bio-oils for the experiments without catalyst and 10.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH were correlated with a 1st order exponential decay and 1st order exponential models, respectively, exhibiting for both models, *root-mean-square error* (r^2) of 0.99, while the experiment with 7.5% (wt.) activated carbon pellets impregnated with 10.0 M NaOH was correlated with a 1st order linear equation, exhibiting *root-mean-square error* (r^2) of 0.92. For the experiment with 7.5% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, the kinematic viscosity was correlated with a *DoseResp* (sigmoid) model, exhibiting *root-mean-square error* (r^2) of

0.98. By *DoseResp* function, A1 and A2 are the bottom (initial value) and top (final value) asymptotes, respectively, being A1 equal to 2.9 and A2 equal to 7.74. The kinematic viscosity of bio-oils decrease because at the beginning of catalytic cracking reaction, the primary cracking occurs, forming oxygenated compounds [14-16, 18]. Then, in the middle to the end of the reaction, secondary cracking occurs by de-carbonylation and de-carboxylation, producing hydrocarbons [14-16, 18, 25-32], and hydrocarbons have much lower viscosity than oxygenates.

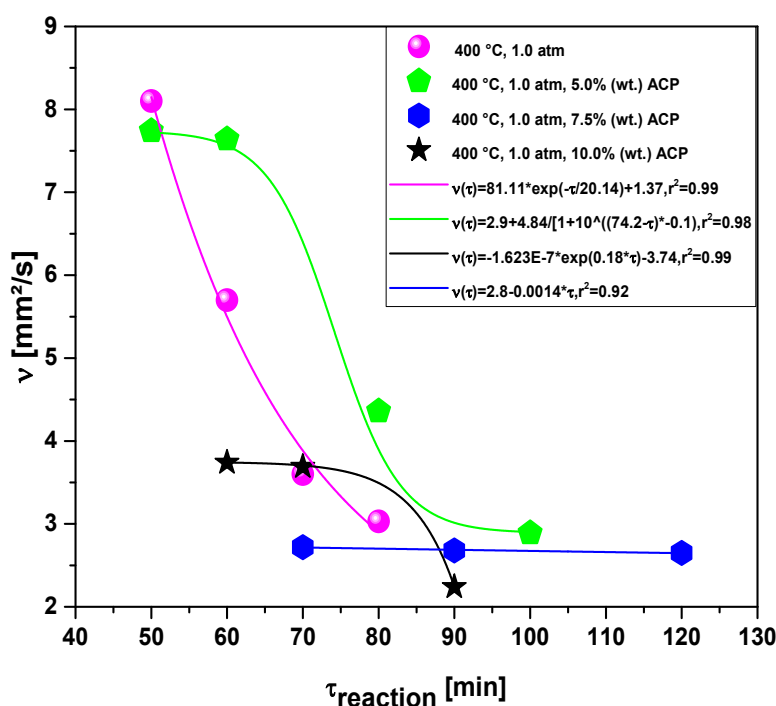


Figure 14. Effect of reaction time on the kinematic viscosity of bio-oil by the upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atmosphere, 0.0, 5.0, 7.5, and 10.0% (wt.) of activated carbon pellets impregnated with 10.0 M NaOH, using a catalyst fixed bed reactor, in a semi-pilot scale two-stage reactor of 2.0 L.

• 4.2.3.3 Effect of reaction time on the acidity of bio-oil

The influence of reaction time on the density of bio-oil by catalytic upgrading of residual fat pyrolysis vapors at 450 °C, 1.0 atm, 0.0, 5.0, 7.5, and 10.0% (wt.) of Red Mud pellets activated with 1.0 M HCl, using a catalyst fixed bed reactor, in semi pilot scale two-stage reactor of 2.0 L, is shown in Figure 15. The acidity of bio-oil, expressed by the acid value, decreases drastically with increasing reaction time. The experiments without catalyst, with 7.5 and 10.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH were correlated with a *first order exponential decay model*, exhibiting root-mean-square error (r^2) between 0.98 and 0.99, while the experiments with 5.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH were correlated with a *first order exponential model*, exhibiting root-mean-square errors (r^2) of 0.99. This is due to the catalytic deoxygenation of triglycerides and fatty acids molecules, the main compounds of vegetable oils and animal fats, into hydrocarbons through de-carboxylation/de-carbonylation as reported in the literature [33-34].

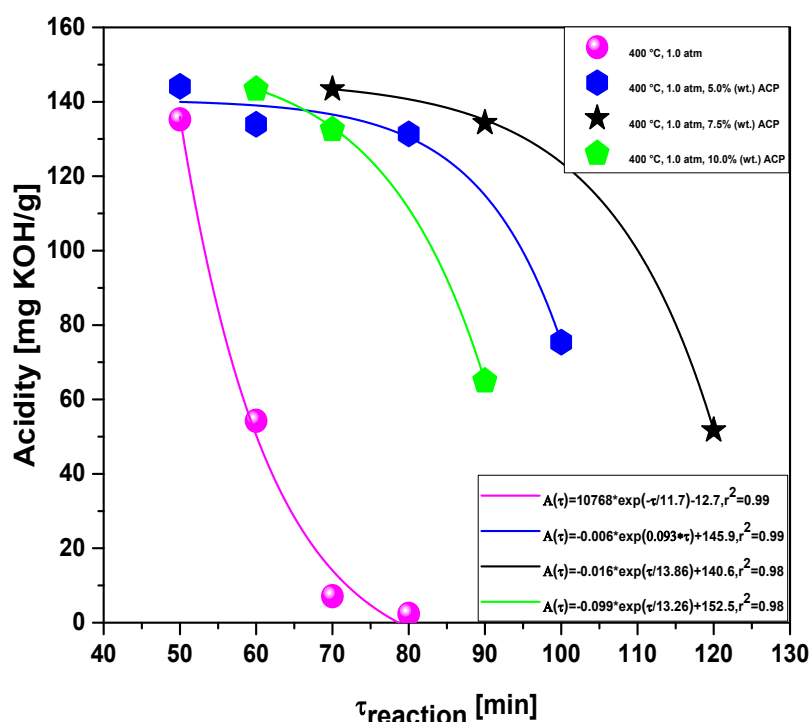


Figure 15. Effect of reaction time on the acidity of bio-oil by the upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atmosphere, 0.0, 5.0, 7.5, and 10.0% (wt.) of activated carbon pellets impregnated with 10.0 M NaOH, using a catalyst fixed bed reactor, using a semi-pilot scale two-stage reactor of 2.0 L.

As observed in Figure 15, as soon as the first sample is withdrawn, the acid value is very high, ranging between 135.26 and 144.14 mg KOH/g, due to formation of carboxylic acids, associated to the primary cracking. Afterwards, in the middle to the end of the reaction, deoxygenation of triglycerides and fatty acids molecules occurs by means of de-carboxylation/de-carbonylation, producing mixtures rich in hydrocarbons and poor of oxygenates, particularly carboxylic acids [33-34].

4.2.4. Effect of reaction time on the selectivity of hydrocarbons and oxygenates in bio-oil

Table 5 and Figures 16 and 17 illustrate the effect of reaction time on the selectivity of hydrocarbons and oxygenates in bio-oils, respectively, by the upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atmosphere, 0.0, 5.0, 7.5, and 10.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, using a catalyst fixed bed reactor, using a semi-pilot scale two-stage reactor of 2.0 L.

The chemical functions (alkanes, alkenes, alkynes, ring-containing alkanes, ring-containing alkenes, aromatics, carboxylic acids, aldehydes, alcohols, and ketones), sum of peak areas, CAS numbers, and retention times of all the molecules identified in bio-oil by GC-MS by the upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atmosphere, 0.0, 5.0, 7.5, and 10.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, during the course of reaction at 50, 60, 70, 80, 90, 100 and 120 min, using a catalyst fixed bed reactor, in a semi-pilot scale two-stage reactor of 2.0 L, are illustrated in Supplementary Tables S6-S19.

Table 5. Effect of reaction time on the concentration of bio-oil by catalytic upgrading of residual fat pyrolysis vapors at 450 °C, 1.0 atm, 0.0, 5.0, 7.5, and 10.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, using a catalyst fixed bed reactor, in a semi-pilot scale two-stage reactor of 2.0 L.

Temperature/Catalyst	t _{Reaction} [min]	Concentration [%area.]	
		Hydrocarbons	Oxygenates
400 °C	50	71.950	28.050
	60	88.515	11.483
	70	90.966	9.034
	80	95.137	4.865
400 °C, 5% ACP (wt.)	50	27.856	72.146
	60	43.619	56.381
	80	48.479	51.524
	100	68.328	31.667
400 °C, 7.5% ACP (wt.)	70	50.836	49.164
	90	55.507	44.493
	120	68.363	31.636
400 °C, 10% ACP (wt.)	60	53.925	46.074
	70	55.290	44.709
	90	75.764	24.237

The concentration of hydrocarbons in bio-oil increases with reaction time. The experiments without catalyst, 7.5 and 10.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH were correlated with a *first order exponential model*, exhibiting for all regressions, root-mean-square errors (r^2) of 0.98, while the experiments with 5.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH was correlated with a 1st order linear equation, exhibiting *root-mean-square error* (r^2) of 0.93. The concentration of hydrocarbons in bio-oil increases due to the catalytic deoxygenation of fatty acids molecules, by means of de-carboxylation/de-carbonylation, producing aliphatic and aromatic hydrocarbons, as reported in the literature [33-34]. The results are according to similar studies reported in the literature [1, 3-4, 6, 8]. On the contrary, the concentration of oxygenates in bio-oil decreases with increasing reaction time. The experiments without catalyst, 7.5 and 10.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH were correlated with a *first order exponential model*, exhibiting for all regressions, root-mean-square errors (r^2) between 0.98 and 0.99, while the experiments with 5.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH was correlated with a 1st order linear equation, exhibiting *root-mean-square error* (r^2) of 0.93.

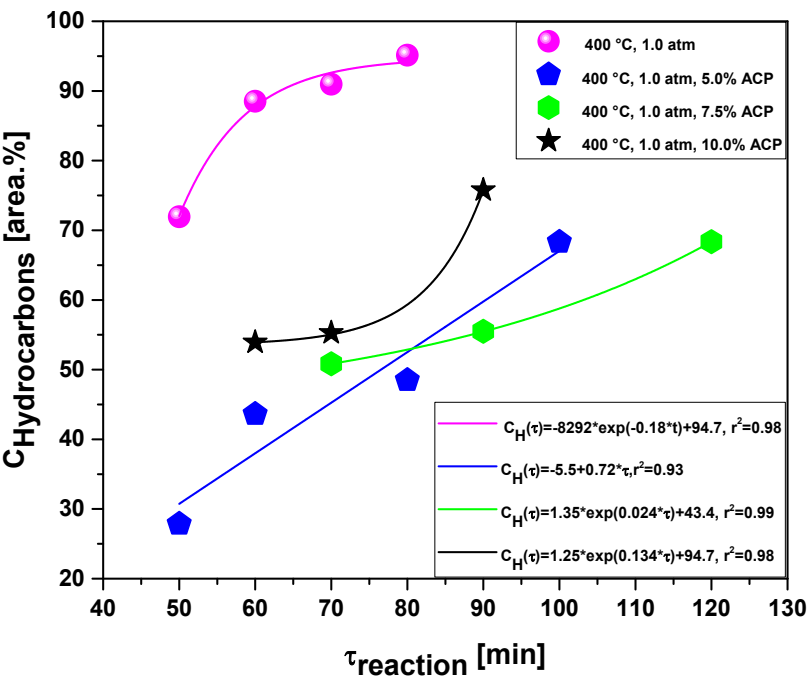


Figure 16. Effect of reaction time on the selectivity of hydrocarbons by the upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atmosphere, 0.0, 5.0, 7.5, and 10.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, using a catalyst fixed bed reactor, in a semi-pilot scale two-stage reactor of 2.0 L.

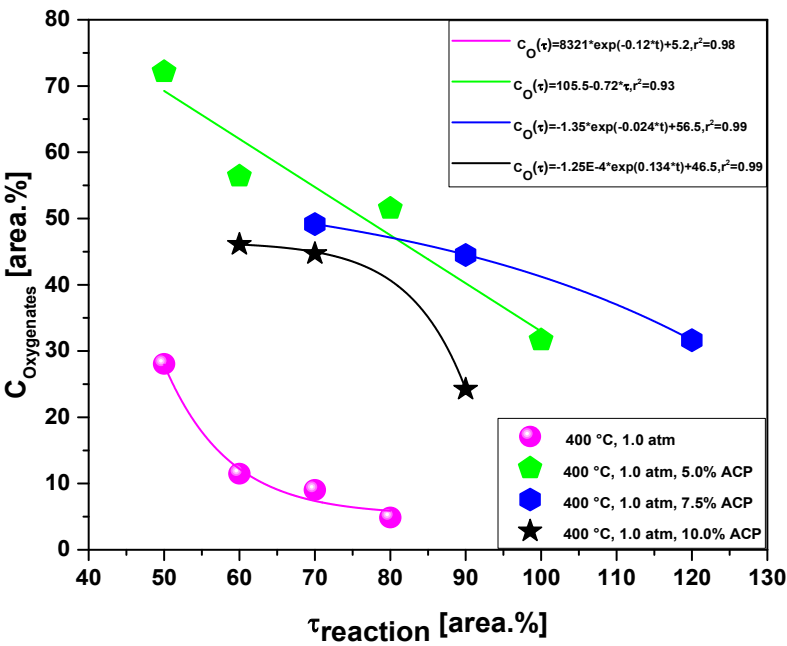


Figure 17. Effect of reaction time on the selectivity of oxygenates by the upgrading of residual fat pyrolysis vapors at 450 °C, 1.0 atmosphere, 0.0, 5.0, 7.5, and 10.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, using a catalyst fixed bed reactor, using a semi-pilot scale two-stage reactor of 2.0 L.

5. Conclusions

The activated carbon pellets impregnated with 10.0 M NaOH were not selective to promote the deoxygenation of triglycerides and fatty acids molecules by means of de-carboxylation/de-carbonylation, as the acidity of bio-oils was high compared to the experiments without catalyst. On the other hand, as the de-carboxylation and de-carbonylation has not taken place, there was little formation of CO and CO₂, is to expect the gaseous products were rich on hydrocarbons of short chain length.

Supplementary Materials: The following are available. ST 6: Classes of compounds, summation of peak areas, CAS number, and retention times of chemical compounds identified by CG-MS in bio-oil by pyrolysis of residual fat at 400 °C, 1.0 atm, 50 min, using a pyrolysis reactor of 2.0 L, in semi pilot scale. ST 7: Classes of compounds, summation of peak areas, CAS number, and retention times of chemical compounds identified by CG-MS in bio-oil by pyrolysis of residual fat at 400 °C, 1.0 atm, 60 min, using a pyrolysis reactor of 2.0 L, in semi pilot scale. ST 8: Classes of compounds, summation of peak areas, CAS number, and retention times of chemical compounds identified by CG-MS in bio-oil by pyrolysis of residual fat at 400 °C, 1.0 atm, 70 min, using a pyrolysis reactor of 2.0 L, in semi pilot scale. ST 9: Classes of compounds, summation of peak areas, CAS number, and retention times of chemical compounds identified by CG-MS in bio-oil by pyrolysis of residual fat at 400 °C, 1.0 atm, 80 min, using a pyrolysis reactor of 2.0 L, in semi pilot scale. ST 10: Classes of compounds, summation of peak areas, CAS number, and retention times of chemical compounds identified by CG-MS in bio-oil by catalytic upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atm, 5.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, 50 min, using a catalyst fixed bed reactor, in semi pilot scale two-stage reactor of 2.0 L. ST 11: Classes of compounds, summation of peak areas, CAS number, and retention times of chemical compounds identified by CG-MS in bio-oil by catalytic upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atm, 5.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, 60 min, using a catalyst fixed bed reactor, in semi pilot scale two-stage reactor of 2.0 L. ST 12: Classes of compounds, summation of peak areas, CAS number, and retention times of chemical compounds identified by CG-MS in bio-oil by catalytic upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atm, 5.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, 80 min, using a catalyst fixed bed reactor, in semi pilot scale two-stage reactor of 2.0 L. ST 13: Classes of compounds, summation of peak areas, CAS number, and retention times of chemical compounds identified by GC-MS in bio-oil by catalytic upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atm, 5.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, 100 min, using a catalyst fixed bed reactor, in semi pilot scale two-stage reactor of 2.0 L. ST 14: Classes of compounds, summation of peak areas, CAS number, and retention times of chemical compounds identified by GC-MS in bio-oil by catalytic upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atm, 7.5% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, 70 min, using a catalyst fixed bed reactor, in semi pilot scale two-stage reactor of 2.0 L. ST 15: Classes of compounds, summation of peak areas, CAS number, and retention times of chemical compounds identified by GC-MS in bio-oil by catalytic upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atm, 7.5% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, 90 min, using a catalyst fixed bed reactor, in semi pilot scale two-stage reactor of 2.0 L. ST 16: Classes of compounds, summation of peak areas, CAS number, and retention times of chemical compounds identified by GC-MS in bio-oil by catalytic upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atm, 7.5% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, 120 min, using a catalyst fixed bed reactor, in semi pilot scale two-stage reactor of 2.0 L. ST 17: Classes of compounds, summation of peak areas, CAS number, and retention times of chemical compounds identified by GC-MS in bio-oil by catalytic upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atm, 10.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, 60 min, using a catalyst fixed bed reactor, in semi pilot scale two-stage reactor of 2.0 L. ST 18: Classes of compounds, summation of peak areas, CAS number, and retention times of chemical compounds identified by GC-MS in bio-oil by catalytic upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atm, 10.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, 70 min, using a catalyst fixed bed reactor, in semi pilot scale two-stage reactor of 2.0 L. ST 19: Classes of compounds, summation of

peak areas, CAS number, and retention times of chemical compounds identified by GC-MS in bio-oil by catalytic upgrading of residual fat pyrolysis vapors at 400 °C, 1.0 atm, 10.0% (wt.) activated carbon pellets impregnated with 10.0 M NaOH, 90 min, using a catalyst fixed bed reactor, in semi pilot scale two-stage reactor of 2.0 L.

Author Contributions: The individual contributions of all the co-authors are provided as follows: L.P.B. contributed with formal analysis and writing original draft preparation, investigation and methodology, C.C.F. contributed with investigation and methodology, A.F.d.F.C. contributed with investigation and methodology, H.J.d.S.R. contributed with investigation and methodology, W.G.d.S. contributed with chemical analysis, L.M.P. contributed with chemical analysis, A.M.P. contributed with chemical analysis, N.L.M. contributed with chemical analysis, F.P.d.C.A. contributed with chemical analysis, S.A.P.d.M. contributed with chemical analysis and formal analysis, D.A.R.d.C. contributed with investigation, methodology and chemical analysis, M.C.S. contributed with formal analysis, investigation and methodology, N.M.M. contributed with contributed with resources, chemical analysis, S.D.J. contributed with resources, chemical analysis, L.E.P.B. with co-supervision, and resources, and N.T.M. contributed with supervision, conceptualization, and data curation. All authors have read and agreed to the published version of the manuscript.

Funding: This research received no external funding.

Institutional Review Board Statement: Not applicable.

Informed Consent Statement: Not applicable.

Acknowledgments: I would like to acknowledge and dedicate this research in memory to Hélio da Silva Almeida, he used to work at the Faculty of Sanitary and Environmental Engineering/UFPa, and passed away on 13 March 2021. His contagious joy, dedication, intelligence, honesty, seriousness, and kindness will always be remembered in our hearts.

Conflicts of Interest: The authors declare no conflict of interest.

References

1. Nur Hidayah Zainan, Srikanth Chakravartula Srivatsa, Fanghua Li, Sankar Bhattacharya. Quality of bio-oil from catalytic pyrolysis of microalgae *Chlorella vulgaris*. *Fuel* 223 (2018) 12–19
2. Yunpu Wang, Qiuhaio Wu, Dengle Duan, Roger Ruan, Yuhuan Liu, Leilei Dai, Yue Zhou, Yunfeng Zhao, Shumei Zhang, Zihong Zeng, Lin Jiang, Zhenting Yu. Ex-situ catalytic upgrading of vapors from fast microwave-assisted copyrolysis of *Chromolaena odorata* and soybean soapstock. *Bioresource Technology* 261 (2018) 306–312
3. Prangtip Kaewpengkrow, Duangduen Atong, Viboon Sricharoenchaikul. Selective catalytic fast pyrolysis of *Jatropha curcas* residue with metal oxide impregnated activated carbon for upgrading bio-oil. *International Journal of Hydrogen Energy* 42 (2017) 18397 – 18409
4. P. Kaewpengkrow, D. Atong, V. Sricharoenchaikul. Catalytic upgrading of pyrolysis vapors from *Jatropha* wastes using alumina, zirconia and titania based catalysts. *Bioresource Technology* 163 (2014) 262–269
5. Linlin Yi, Huan Liu, Meiyong Li, Gaozhi Man, Hong Yao. Prevention of CaO deactivation using organic calcium precursor during multicyclic catalytic upgrading of bio-oil. *Fuel* 271 (2020) 117692
6. Supawan Vichaphund, Duangdao Aht-ong, Viboon Sricharoenchaikul, Duangduen Atong. Catalytic upgrading pyrolysis vapors of *Jatropha* waste using metal promoted ZSM-5 catalysts: An analytical PY-GC/MS. *Renewable Energy* 65 (2014) 70-77
7. Yunpu Wang, Qiuhaio Wu, Sha Yang, Qi Yang, Junlin Wu, Zhiyun Ma, Lin Jiang, Zhenting Yu, Leilei Dai, Yuhuan Liu, Roger Ruan, Guiming Fu, Bo Zhang, Haibin Zhu. Microwave-assisted catalytic fast pyrolysis coupled with microwave absorbent of soapstock for bio-oil in a downdraft reactor. *Energy Conversion and Management* 185 (2019) 11–20
8. Qiuhaio Wu, Yunpu Wang, Lin Jiang, Qi Yang, Linyao Ke, Yujie Peng, Sha Yang, Leilei Dai, Yuhuan Liu, Roger Ruan. Microwave-assisted catalytic upgrading of co-pyrolysis vapor using HZSM-5 and MCM-41 for bio-oil production: Co-feeding of soapstock and straw in a downdraft reactor. *Bioresource Technology* 299 (2020) 122611
9. Qiuhaio Wu, Yunpu Wang, Yujie Peng, Linyao Ke, Qi Yang, Lin Jiang, Leilei Dai, Yuhuan Liu, Roger Ruan, Donghua Xia, Li Jiang. Microwave-assisted pyrolysis of waste cooking oil for hydrocarbon bio-oil over metal oxides and HZSM-5 catalysts. *Energy Conversion and Management* 220 (2020) 113-124
10. Wenchao Ma, Bin Liu, Ruixue Zhang, Tianbao Gu, Xiang Ji, Lei Zhong, Guanyi Chend, Longlong Ma, Zhanjun Cheng, Xiangping Li. Co-upgrading of raw bio-oil with kitchen waste oil through fluid catalytic cracking (FCC). *Applied Energy* 217 (2018) 233–240

11. Josué Alves Melo, Mirele Santana de Sá, Ainara Moral, Fernando Bimbela, Luis M. Gandía and Alberto Wisniewski, Jr. Renewable Hydrocarbon Production from Waste Cottonseed Oil Pyrolysis and Catalytic Upgrading of Vapors with Mo-Co and Mo-Ni Catalysts Supported on γ -Al₂O₃. *Nanomaterials* 2021, 11, 1659
12. Zhenting Yu, Lin Jiang, Yunpu Wang, Yanzhi Li, Linyao Ke, Qi Yang, Yujie Peng, Jiamin Xu, Leilei Dai, Qiuha Wu, Yuhuan Liu, Roger Ruan, Donghua Xia, Li Jiang. Catalytic pyrolysis of woody oil over SiC foam-MCM41 catalyst for aromatic-rich bio-oil production in a dual microwave system. *Journal of Cleaner Production* 255 (2020) 120179
13. Mithelesh Koul, Krushna Prasad Shadangi, Kaustubha Mohanty. Effect of catalytic vapour cracking on fuel properties and composition of castor seed pyrolytic oil. *Journal of Analytical and Applied Pyrolysis* 120 (2016) 103–109
14. H. da Silva Almeida, O.A. Corrêa, J.G. Eid, H.J. Ribeiro, D.A.R. de Castro, M.S. Pereira, L.M. Pereira, A. de Andrade Mâncio, M.C. Santos, J.A. da Silva Souza, Luiz E.P. Borges, N.M. Mendonça, N.T. Machado. Production of biofuels by thermal catalytic cracking of scum from grease traps in pilot scale. *Journal of Analytical and Applied Pyrolysis* 118 (2016) 20–33
15. H. da Silva Almeida, O.A. Corrêa, J.G. Eid, H.J. Ribeiro, D.A.R. de Castro, M.S. Pereira, L.M. Pereira, A. de Andrade Mâncio, M.C. Santos, S.A.P. da Mota, J.A. da Silva Souza, Luiz E.P. Borges, N.M. Mendonça, N.T. Machado. Performance of thermochemical conversion of fat, oils, and grease into kerosene-like hydrocarbons in different production scales. *Journal of Analytical and Applied Pyrolysis* 120 (2016) 126–143
16. H. da Silva Almeida, O.A. Corrêa, C.C. Ferreira, H.J. Ribeiro, D.A.R. de Castro, M.S. Pereira, A. de Andrade Mâncio, M.C. Santos, S.A.P. da Mota, J.A. da Silva Souza, Luiz E.P. Borges, N.M. Mendonça, N.T. Machado. Diesel-like hydrocarbon fuels by catalytic cracking of fat, oils, and grease (FOG) from grease traps. *Journal of the Energy Institute* 90 (2017) 337–354
17. Douglas Alberto Rocha de Castro, Haroldo Jorge da Silva Ribeiro, Lauro Henrique Hamoy Guerreiro, Lucas Pinto Bernar, Sami Jonatan Bremer, Marcelo Costa Santo, Hélio da Silva Almeida, Sergio Duvoisin, Jr., Luiz Eduardo Pizarro Borges, Nélito Teixeira Machado. Production of Fuel-Like Fractions by Fractional Distillation of Bio-Oil from Açaí (*Euterpe oleracea* Mart.) Seeds Pyrolysis. *Energies* 2021, 14, 3713
18. S.A.P. da Mota, A.A. Mancio, D.E.L. Lhamas, D.H. de Abreu, M.S. da Silva, W.G. dos Santos, D.A.R. de Castro, R.M. de Oliveira, M.E. Araújo, Luiz E.P. Borges, N.T. Machado. Production of green diesel by thermal catalytic cracking of crude palm oil (*Elaeis guineensis* Jacq) in a pilot plant. *Journal of Analytical and Applied Pyrolysis* 110 (2014) 1–11
19. Amrit Kumar Thakur, Ravishankar Sathyamurthy, R. Velraj, I. Lynch, R. Saidur, A. K. Pandey, Swellam W. Sharshir, Zhenjun Ma, P. GaneshKumar, A.E. Kabeel. Sea-water desalination using a desalting unit integrated with a parabolic trough collector and activated carbon pellets as energy storage medium. *Desalination* 516 (2021) 115217
20. W. G. Appleby, J. W. Gibson, and G. M. Good. *Industrial & Engineering Chemistry Process Design and Development* 1962 1 (2), 102–110, DOI: 10.1021/i260002a006
21. Eduardo Santillan-Jimenez and Mark Crocker. Catalytic deoxygenation of fatty acids and their derivatives to hydrocarbon fuels via decarboxylation/decarbonylation. *J Chem Technol Biotechnol* 2012;87: 1041 – 1050
22. Meriem Fertani-Gmati, Khemaies Brahim, Ismail Khattech, Mohamed Jemal. Thermochemistry and kinetics of silica dissolution in NaOH solutions: Effect of the alkali concentration. *Thermochimica Acta* 594 (2014) 58–67
23. Marit Beerse, Katrien Keune, Piet Iedema, Sander Woutersen, and Joen Hermans. Evolution of Zinc Carboxylate Species in Oil Paint Ionomers. *ACS Appl. Polym. Mater.* 2020, 2, 5674–5685
24. Dengle Duan, Yunpu Wang, Leilei Dai, Roger Ruan, Yunfeng Zhao, Liangliang Fan, Maimaitiaili Tayier, Yuhuan Liu. Ex-situ catalytic co-pyrolysis of lignin and polypropylene to upgrade bio-oil quality by microwave heating. *Bioresour Technol* 241 (2017) 207–213
25. Alisson Caio Magalhães Serrão, Conceição Maria Sales Silva, Fernanda Paula da Costa Assunção, Haroldo Jorge da Silva Ribeiro, Marcelo Costa Santos, Hélio da Silva Almeida, Sergio Duvoisin Jr., Luiz Eduardo Pizarro Borges, Douglas Alberto Rocha de Castro, Nélito Teixeira Machado. Process analysis of pyrolysis of Açaí (*Euterpe Oleracea*, Mart) seeds: Influence of temperature on the yield of reaction products and physico-chemical properties of Bio-Oil. *Brazilian Journal of Development*, Curitiba, v.7, n.2, p. 18200–18220 feb. 2021
26. Mota, S. A. P., Mancio, A. A., Santanna, J. S., Gama, V. de J. P., & Machado, N. T. (2021). Influência do tempo de reação sobre as propriedades físico-químicas de biocombustíveis obtidos através do craqueamento térmico catalítico do óleo de palma bruto. *Scientia Plena*, 17(6). <https://doi.org/10.14808/sci.plena.2021.064201>
27. Flávio Pinheiro Valois, Ana Rosa Bezerra Cardoso, Raymundo da Costa França Neto, Lia Martins Pereira, Dyenny Ellen Lima Lhamas, Sílvio Alex Pereira da Mota, Hélio da Silva Almeida, Luiz Eduardo Pizarro Borges, Nélito Teixeira Machado, Marcelo Costa Santos. Thermal catalytic cracking of soap phase residue of neutralization process of palm oil using CaCO₃ as catalyst. *Brazilian Journal of Development*, Curitiba, v.7, n.6, p.59461–59481 jun.2021
28. Romero Moreira de Oliveira, Emerson Cardoso Rodrigues, Dilson Nazareno Pereira Cardoso, Wenderson Gomes dos Santos, Nélito Teixeira Machado. Thermocatalytic cracking of fat from fat boxes with activated red mud. *Braz. J. of Develop.*, Curitiba, v. 6, n.4, p.19876–19887 apr.2020. ISSN 2525–8761
29. Lia Martins Pereira, Nélito Teixeira Machado, Hélio da Silva Almeida, Fernanda Paula da Costa Assunção, Laércio dos Santos Rosa Junior. Characterization of biocarvia through catalytic thermal cracking from blend of sewage and residual fat in scale pilot. *Braz. J. of Develop.*, Curitiba, v. 6, n. 1, p.1502–1509 jan. 2020. ISSN 2525–8761

-
30. A.A. Mancio, S.A.P. da Mota, C.C. Ferreira, T.U.S. Carvalho, O.S. Neto, J.R. Zamian, M.E. Araújo, L.E.P. Borges, N.T. Machado. Separation and characterization of biofuels in the jet fuel and diesel fuel ranges by fractional distillation of organic liquid products. *Fuel* 215 (2018) 212–225
 31. A.A. Mancio, K.M.B. da Costa, C.C. Ferreira, M.C. Santos, D.E.L. Lhamas, S.A.P. da Mota, R.A.C. Leão, R.O.M.A. de Souza, M.E. Araújo, L.E.P. Borges, N.T. Machado. Process analysis of physicochemical properties and chemical composition of organic liquid products obtained by thermochemical conversion of palm oil. *Journal of Analytical and Applied Pyrolysis* 123 (2017) 284–295
 32. M.C. Santos, R.M. Lourenço, D.H. de Abreu, A.M. Pereira, D.A.R. de Castro, M.S. Pereira, H.S. Almeida, A.A. Mancio, D.E.L. Lhamas, S.A.P. da Mota, J.A. da Silva Souza, S.D. Júnior, M.E. Araújo, L.E.P. Borges, N.T. Machado. Gasoline-like hydrocarbons by catalytic cracking of soap phase residue of neutralization process of palm oil (*Elaeis guineensis* Jacq). *Journal of the Taiwan Institute of Chemical Engineers* 71 (2017) 106–119
 33. Eduardo Santillan-Jimenez, Tonya Morgan, Joseph Lacny, Susanta Mohapatra, Mark Crocker. Catalytic deoxygenation of triglycerides and fatty acids to hydrocarbons over carbon-supported nickel. *Fuel* 103 (2013) 1010–1017
 34. Hao Chena, Yulong Wub, Suitao Qi, Yu Chenb, Mingde Yang. Deoxygenation of octanoic acid catalyzed by hollow spherical Ni/ZrO₂. *Applied Catalysis A: General* 529 (2017) 79–90