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Review

The Technologies of Electrochemical Extraction Lithium Process from Lithium-Containing Solutions

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Abstract: With the rapid development of new energy vehicles and the digital electronics industry, the demand for lithium has surged, necessitating advanced lithium extraction technologies. Electrochemical methods, noted for their high selectivity and efficiency in extracting target ions from liquid sources in an environmentally friendly manner, have become increasingly vital. These methods are versatile, applicable in scenarios such as lithium extraction from saline lakes, mother liquor separation, and lithium enrichment. They include electrochemical deintercalation, electrochemical ion pumps, and electrodialysis, each offering unique benefits and challenges depending on the application context. This review provides a detailed exploration of the research progress in lithium extraction using electrochemical methods and discusses future prospects for these technologies, emphasizing their potential to meet the growing demand for lithium.

Keywords: lithium extraction; electrochemistry; waste recycling; environment protection

1. Introduction

With the continuous advancement of the new energy sector, lithium-ion batteries, which offer advantages such as long service life and mature technology, have been applied in various aspects of daily life [1–3]. Particularly in recent years, following the rapid growth of the new energy sector and the continuous expansion of the electric vehicle market, the consumption of lithium resources has been escalating annually [4–6]. In light of the Statistical Multifragmentational Model's (SMM), the worldwide demand for lithium batteries witnessed a year-on-year growth of 52% in 2022, while the supply of lithium resources reached as high as 840 Kt [7]. From a global perspective, a notable mismatch in the regional supply and demand of lithium resources is evident. Lithium ore exploitation resources in the upstream of the supply chain are mainly distributed in Australia, Chile, Argentina, and other regions in South America [8,9]. Thus, the extraction of lithium resources from secondary resource lithium becomes the focus of researchers. Secondary lithium resources are generated along with the slag from lithium ore [10], hydrometallurgical leaching the lithium from waste lithium batteries [11]. With the continuous expansion of the scale of lithium ore production and the recycling of waste lithium-ion batteries, the total amount of lithium-rich solutions shows an increasing trend year by year [12,13]. Therefore, the efficient and convenient lithium extraction technology is worthy study in the future.

Currently, a wide array of lithium extraction technologies are employed to recover lithium from lithium-containing solutions, including hydrometallurgy [14,15] and electrochemical recovery processes [16,17]. The main technology is hydrometallurgy method, through the co-precipitation method to recover the metal elements in the lithium solution, and reproduced as chemical raw materials to achieve the recycling of lithium resources. Electrochemical extraction technology is an efficient lithium extraction technology that has rapidly developed in recent years [17,18]. The primary

technology used is the hydrometallurgy method, which utilizes co-precipitation to recover metal elements from lithium solutions and reproduce them as chemical raw materials, thereby achieving the recycling of lithium resources [19–21]. The electrochemical lithium extraction technology features high efficiency, environmental friendliness, and low energy consumption, with extensive prospects in the future [22,23]. This review aims to analyze the properties of various electrochemical lithium extraction methods from lithium-containing solutions, outline the current research status and technical characteristics, and promote the development of more efficient and environmentally friendly electrochemical lithium extraction technologies to meet the increasing demand for lithium extraction.

2. Different techniques of electrochemical lithium extraction

2.1. Electrochemical Deintercalation Method

LiFePO₄ is a commonly used anode materials of lithium ion battery. LiFePO₄ has a typical olivine-type crystal structure [23,24]. The crystal structure of LiFePO₄ is composed of [LiO₆] octahedra and [PO₄] tetrahedra, which facilitate the migration of lithium ions and electronic conduction through the Fe-O-Fe lattice planes. The chemical reactions involved in the removal and insertion of Li⁺ in the LiFePO₄ crystal structure are depicted in equations (1) and (2) [25] (Figure 1).



Based on above electrochemical delithiation theory, lithium iron phosphate is widely used for lithium extraction and lithium enrichment. Zhao et al. [26–29] have conducted extensive research on lithium extraction using the LiFePO₄ electrode. They developed a "rocking chair" electrode system, which alternates between a lithium-rich LiFePO₄ electrode and a lithium-poor Li_(1-x)FePO₄ electrode. This system is divided by an anion-exchange membrane, with the electrolytic cell containing initial liquid on one side and lithium-rich liquid on the other. Under a direct current electric field, the lithiation reaction occurred at the lithium-poor electrode (FePO₄ cathode) where lithium ions were absorbed from initial liquid, while the delithiation reaction occurred at the lithium-rich electrode (LiFePO₄ anode), diffusing lithium ions into the lithium-rich liquid. Anion exchange membranes are used, allowing Cl⁻ ions to migrate from the cathode to the anode side to maintain charge balance. By switching the positions of the two electrodes and repeating the process, selective enrichment of lithium ions is achieved. Experimental results indicated that the maximum amount of lithium embedded reaches 38.93 mg/g after 10 h, while the maximum amount of embedded Mg²⁺ is 5 mg/g in MgCl₂ solution (220 mg/L). The selective enrichment of lithium ions has been confirmed in a simulated brine composed of LiCl and MgCl₂ solutions. The initial concentration of Li⁺ was 220 mg/L, and the final Mg/Li ratio decreased from 60 to 45, demonstrating effective lithium ion selectivity in the process..

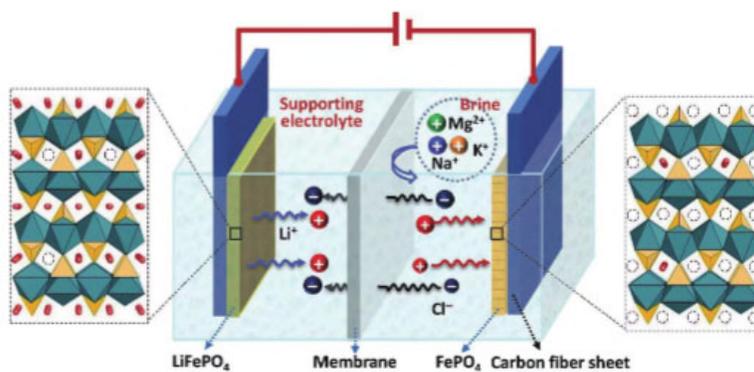


Figure 1. Structure of the LiFePO₄/FePO₄ electrolytic cell for lithium extraction [27].

Deng et al [30] carried out electrochemical extraction lithium using a $\text{LiFePO}_4/\text{FePO}_4$ electrochemical system. The electrode was modified using PEG-6000 as a poregen to obtain abundant pore structures of about 5–20 μm formed on the surface of electrode. This modification increased the adsorption capacity to 17.10 mg g^{-1} , a 57.02% improvement. In this electrochemical system, delithium occurred at the LiFePO_4 electrode, while the lithiation reaction took place at the FePO_4 electrode. The system was effectively employed to recover lithium resources from geothermal water with low lithium concentration. When the modified LiFePO_4 and FePO_4 electrodes were utilized in an H-type electrolytic cell with operated voltage of 1 V, the retrieval rate of Li^+ reached 90.65% after 8 cycles in the geothermal water, and the electrode structure of both anode and cathode maintained stable (Figure 2).

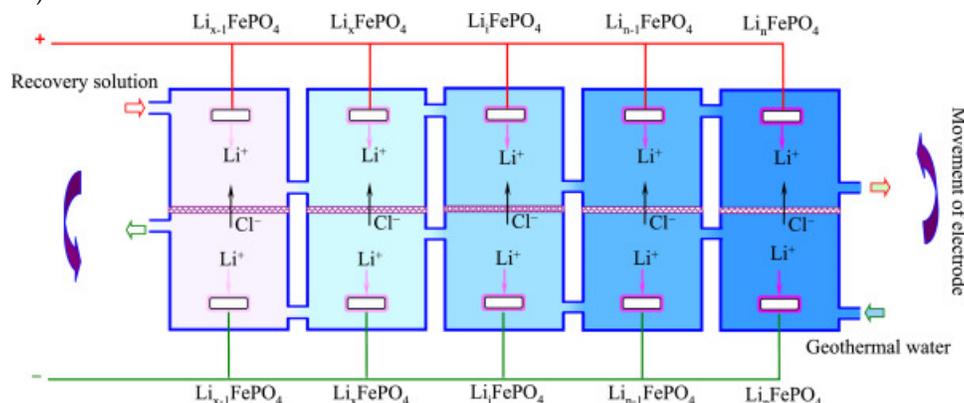


Figure 2. The schematic diagram of continuous electrochemical device proposed for Li^+ recovery from geothermal water [30].

The electrochemical deintercalation method is commonly utilized for lithium extraction. Researchers have observed that the diffusion of anions is driven by the concentration differences of cations between the anode and cathode, leading anions to migrate towards the opposite electrode [31]. As a result, a small number of anions and impurity ions pass through anion-exchange membrane under the driving force of concentration difference, and the separation rate of Li^+ is declined. Specifically, in the treatment of Li_2SO_4 solutions, the relatively large ionic volume of SO_4^{2-} increases migration resistance. This not only necessitates higher voltages for the electrochemical adsorption reaction, prompting numerous side reactions, but also results in substantial electrical resistance and high energy consumption during the process [32].

While the electrochemical deintercalation method is straightforward and easy to operate, it has several drawbacks that require improvement. Specifically, the ionic conduction ability of the membrane and the stability of the electrode material need significant enhancements to optimize the process.

2.2. Electrochemical ion pump

The principle of electrochemical ion pump involves extraction lithium under inverse concentration gradient. The Lithium ions are captured selectively by lithium absorption materials, while the counter electrodes (such as Pt and Ag) are adopted to capture anions [33–35]. Initially reported in the early 1990s by Kanoh and his team, the electrochemical ion pump method utilized $\lambda\text{-MnO}_2$ as the working electrode and a Pt wire as the counter electrode to extract lithium ions from a polycationic solution [36,37]. Subsequently, La Mantia and Yoon's group innovated on this by replacing the Pt electrode with an Ag electrode, and its working principle is shown in Figure 3. The $\lambda\text{-MnO}_2$ serves as the intercalating electrode, adsorbing and desorbing lithium ions, while the Ag electrode functions as the counter electrode, capturing and releasing chloride ions,, thereby significantly reducing the power consumption during the electrochemical reaction [38].

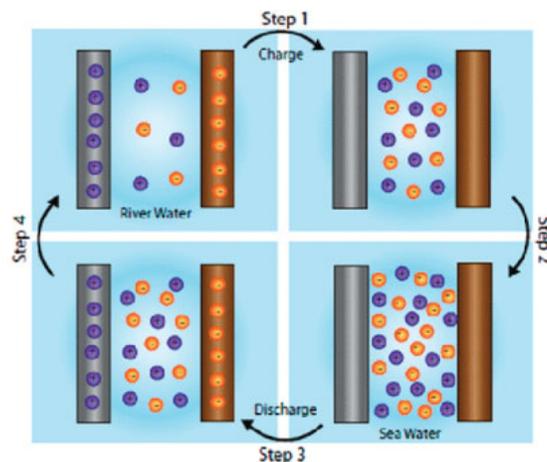


Figure 3. The schematic diagram of electrochemical extraction lithium by ion pump [33].

In recent years, with continuous advancements in materials manufacturing technology, the scope of electrochemical ion pumping applications has notably expanded. Deng et al introduced a green and innovative $\text{FePO}_4/\text{K}_2\text{NiFe}(\text{CN})_6$ [FPO/KNiFC] electrochemical ion pump technique to recover lithium from low-concentration sources in geothermal water (Figure 4). The performance of the LiFePO_4 electrode was enhanced by modifying it with multiwalled carbon nanotubes, improving both the insertion capacity and rate. Although the single insertion capacity stands at 14.88 mg/g, the recovery rate can be increased through cycle insertion-release processes [39]. Moreover, Kazuya Sasaki et al. developed a new electrochemical pumping technology for the extraction/recovery of lithium. This system comprises two power supplies, three electrodes, and a $\text{La}_{0.57}\text{Li}_{0.29}\text{TiO}_3$ (LLTO) electrolyte [40] (Figure 5). The distinctive structure enables the positive voltage of the third electrodes in the cathode solution to be applied to the conventional battery through a secondary power supply. LLTO serves as an ideal solid electrolyte for high-purity lithium extraction and recovery, effectively blocking the permeation of cations other than lithium ions. Notably, this newly designed system demonstrates flexibility in adapting to a variety of commercial electrolytes and capability of significantly enhancing the lithium collection rate of the mentioned main circuit [41]. Furthermore, an additional subcircuit provides effective control, accelerating the lithium absorption process by at least 464 times compared to traditional systems.

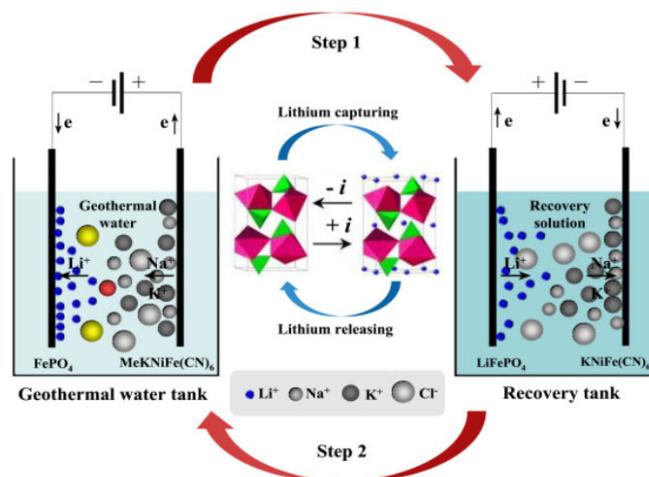


Figure 4. Apparatus and mechanism for lithium recovery from geothermal water by the FPO/KNiFC ion pump technique [39].

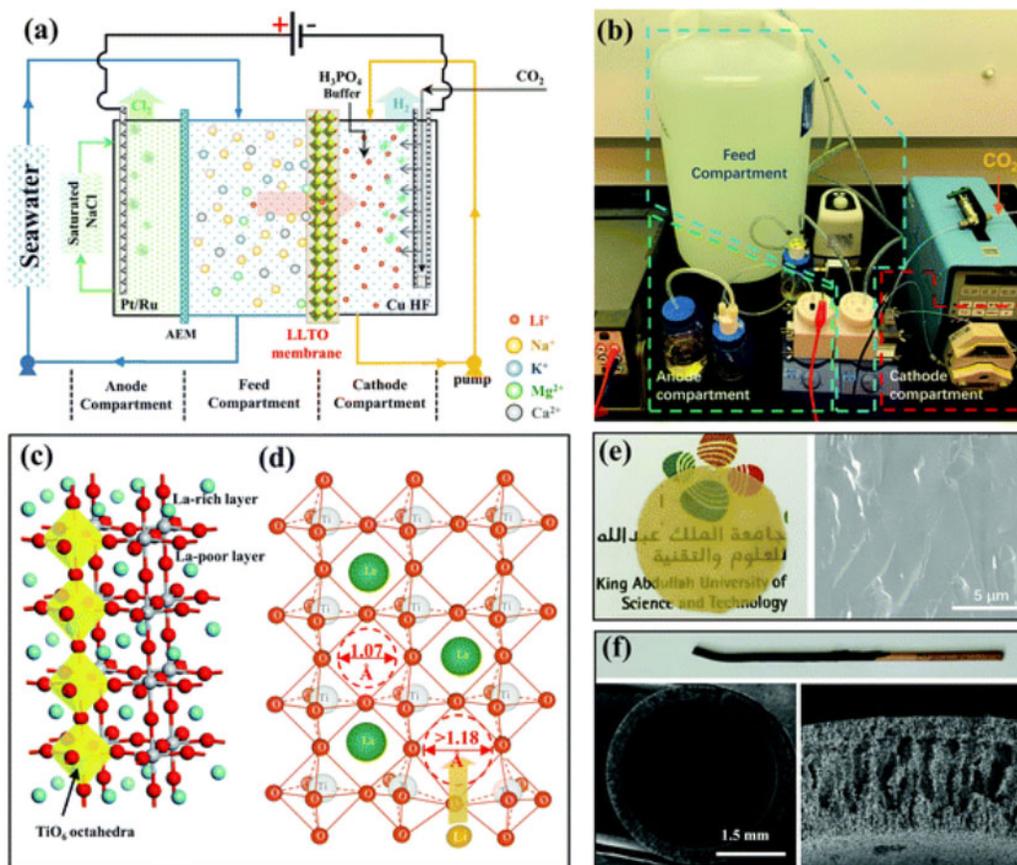


Figure 5. A Schematic of newly designed electrochemical pumping cell comprising a $\text{La}_{0.57}\text{Li}_{0.29}\text{TiO}_3$ (LLTO) electrolyte, Pt anode and cathode, a third Ni electrode, and main and secondary power supplies, thereby enabling mass transfer during Li ion extraction/recovery [41].

Ion pump lithium-extraction technology surpasses other electrochemical methods in selectivity, cycling stability, and energy efficiency, offering significant advantages for practical applications. The success of this technology hinges on electrode materials that exhibit high selectivity and a robust ion exchange capacity. Currently, the use of ion pump technology for lithium recovery remains predominantly within the realm of laboratory research. For it to transition to industrial applications, extensive further research is essential to address practical implementation challenges and scale-up processes.

2.3. Electrodialysis method

Electrodialysis is an electrochemistry separation technology based on selectivity membrane, which was originally developed in the field of seawater desalination [42–44]. And with the development of membrane materials, electrodialysis began to be applied for lithium recovery from saline and other lithium solution. The principle of electrodialysis technology is that a voltage is applied between the two electrodes, driving ions migration across the membranes to achieve the enrichment of target ions [45,46]. The working principle of the electrodialysis system is schemed in Figure 6. The electrodialysis system consists of repetitive units including cation-exchange membrane (CEM) and an anion-exchange membrane (AEM), which separate concentration and dilution chambers alternately. The monovalent ions (e.g., Li^+ , Na^+ , K^+ , and Cl^-) could permeated through these chambers and pass through the CEM and the AEM into the concentration chamber, while the divalent ions (e.g., Mg^{2+} , Ca^{2+} , SO_4^{2-}) are blocked by the ion exchange membrane and retained in the desalination chamber.

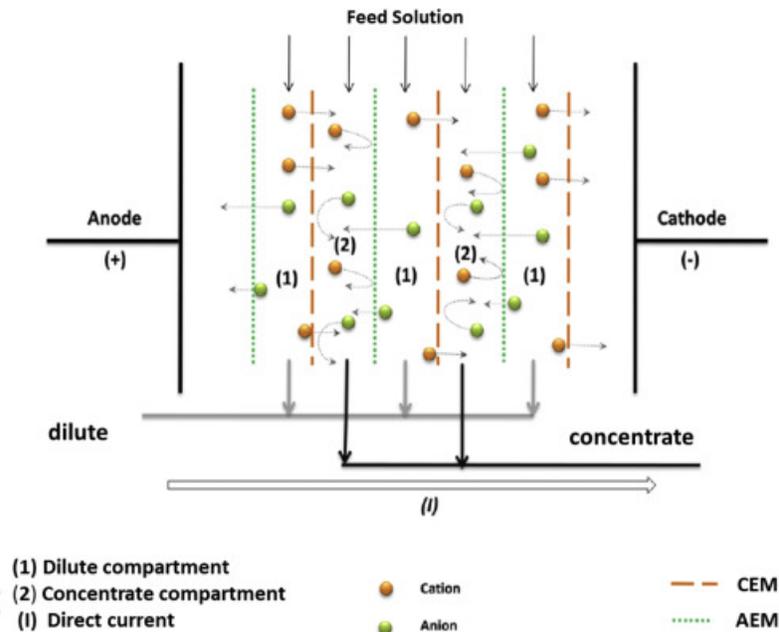


Figure 6. The schematic of electrodesialysis [45].

Abuliti Abudula et. al. developed a novel lithium-ion separation composite membrane composed of $H_{1.6}Mn_{1.6}O_4$ nanoparticles and reduced graphene oxide (rGO) by vacuum filtration technology [47]. The membrane's effectiveness in selectively extracting low-concentration lithium ions from aqueous solutions is attributed to the selective adsorption properties of $H_{1.6}Mn_{1.6}O_4$ and the excellent conductivity of rGO (Figure 7). The results showed that the lithium ion adsorption capacity of $H_{1.6}Mn_{1.6}O_4$ /rGO composite membranes reached an equilibrium of 38.78 mg/g after 5 h. Impressively, the adsorption capacity of these composite membranes was maintained at 99% even after 5 cycles.. Specifically, during selective adsorption experiments with initial molar ratios of Li^+/Na^+ and Li^+/Mg^{2+} set at 1:1, the adsorption ratios achieved were 10.39 for Li^+/Na^+ and 10.23 for Li^+/Mg^{2+} , respectively, demonstrating significant selectivity in ion exchange.

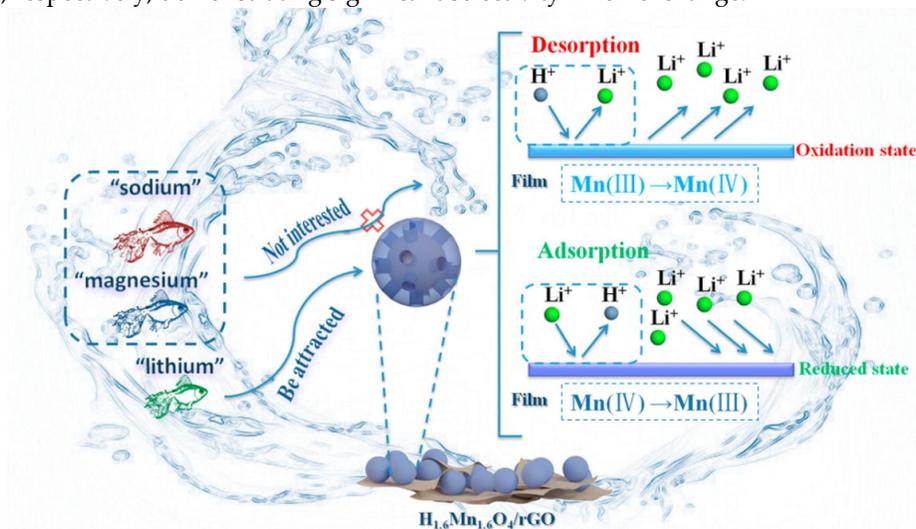


Figure 7. Schematic of the adsorption and desorption on the $H_{1.6}Mn_{1.6}O_4$ /rGO composite film [47].

Shao et al [48] developed a multistage selective electrodesialysis (S-ED) process utilizing high-performance monovalent cation exchange membranes (MCEMs) as shown in Figure 8. These MCEMs were tailored to treat simulated osmotic water produced by a one-stage nanofiltration (NF) process.

The mussel-inspired gallic acid/polyethyleneimine assembled membrane (M-GA/PEI) its unique specific architecture and charge properties demonstrated better separation performance than CSO (a commercial MCEM) during the multi-stage S-ED process. Leveraging its specific architectural and charge characteristics, the M-GA/PEI achieved a Li^+ -rich solution with a high concentration (approximately 8.33 g/L) and purity over 96.4% simultaneously, under a volume ratio of 10:1 (under DC:CC converter) during both the first and second-stage S-ED processes.

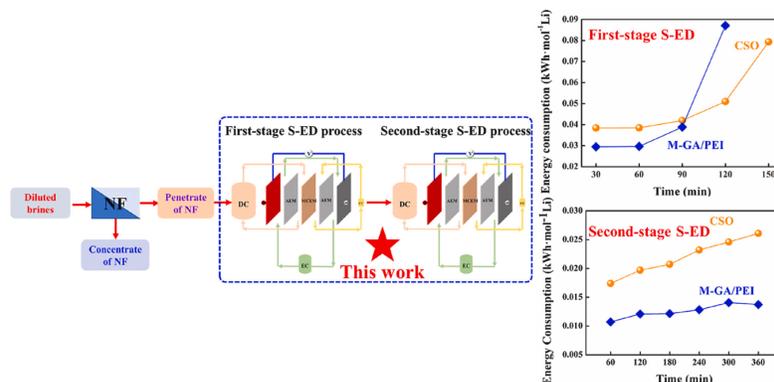


Figure 8. The configuration of the designed selective electrodesialysis [48].

Liu et al introduced an innovative electrodesialysis technology which utilized $\text{Li}_{1.3}\text{Al}_{0.3}\text{Ti}_{1.7}(\text{PO}_4)_3$ (LATP), a lithium superionic conductors, and incorporates an assisted charge balance (ACB) system [49]. The electrodesialysis technology was operated all under low voltage and ultra-low current conditions (Figure 9). The excellent ion selectivity of LATP, coupled with the improved recovery of ACB, enable this technology to achieve exceptionally high separation efficiencies and significantly reduced energy consumption. By implementing a two-stage extraction process on aged brine, the application of directly producing battery-grade Li_2CO_3 with a purity of 99.93%. Remarkably, when applied to pristine brine,, the technology achieves an impressive Li/Mg separation coefficient as high as 5924. This simplified production process is expected to significantly reduce the cost of Li_2CO_3 production.

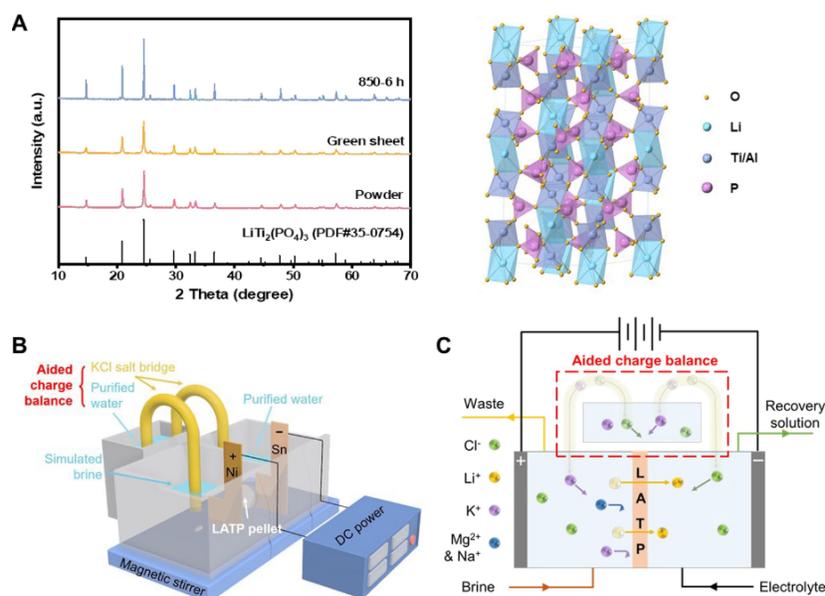


Figure 9. (A) XRD pattern and the crystal structure of LATP, and the schematic illustration of (B) LES and (C) the extraction process [49].

The electro dialysis method offers significant advantages in terms of cost-efficiency, environmental sustainability, and effective separation of lithium ions from bivalent cations. However, its effectiveness in separating lithium from sodium ions is limited. Additionally, the presence of specific anions in the solution, particularly high concentrations of SO_4^{2-} , can dramatically reduce the lifespan of the electro dialysis membranes [50,51]. Another challenge with this method is the generation of substantial amounts of wastewater that require further treatment. Moreover, the migration rate of lithium ions during electro dialysis is relatively slow, leading to lengthy processing times. Consequently, while electro dialysis presents potential benefits, its application in lithium extraction requires more in-depth study to overcome these operational challenges.

3. Conclusions

Electrochemical technology holds promising prospects in lithium extraction and adsorption, encompassing methods such as electrochemical deintercalation, ion pumps, and electro dialysis. These techniques offer advantages such as low cost, environmental sustainability, and reusability. Despite their potential, many of these technologies are still under development, facing significant barriers to commercialization. To advance the field of electrochemical lithium extraction, further research should focus on the following aspects:

(1) Electrochemical materials play a crucial role in determining the performance of lithium extraction systems. The effectiveness of both the anode and cathode materials in the electrochemical system, the ion exchange active materials in electrically controlled setups, and the membrane materials in electro dialysis processes directly influences the efficiency of lithium extraction. Therefore, there is a pressing need to enhance basic research into the development of new materials. Challenges such as low ionic mobility and suboptimal cycling stability persist in current materials. Addressing these issues could involve doping and coating materials, as well as engineering electrodes with perforations to modify their microstructures and chemical compositions, thereby improving their functional properties. In electro dialysis, the choice of membrane materials is critical for effectively separating monovalent from polyvalent ions. Advancements in this area should focus on increasing the ion conductivity, minimizing electrochemical impedance, and enhancing the chemical stability of these membranes. Such improvements will not only boost the performance of lithium extraction but also enhance the overall sustainability and efficiency of the electrochemical methods employed.

(2) Optimization of electrochemical processes. In the past, the industrial production of lithium primarily focused on generating industrial-grade Li_2CO_3 . However, advances in electrochemical production technology are now paving the way for the direct one-step production of battery-grade lithium compounds such as Li_2CO_3 , LiOH , and LiCl . This shift enables not only a transition from single to multiple electrochemical processes but also a transformation of the industry towards refining chemicals and achieving battery-grade production through process optimization. Electrochemical lithium extraction optimally utilizes lithium-containing solutions and recovers lithium from low-concentration and economically unviable brines efficiently. As global efforts to reduce carbon emissions continue to evolve, industrial practices must adapt to minimize environmental impact, carbon emissions, energy, and freshwater consumption. Upgrades in electrochemical lithium extraction technology are essential for significantly improving these parameters, which in turn will accelerate the rapid industrialization and commercialization of lithium extraction processes. This evolution is critical for supporting the growing demand for lithium in various high-tech applications, including electric vehicles and renewable energy storage solutions.

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References

1. Chen, Y.; Kang, Y.; Zhao, Y.; Wang, L.; Liu, J.; Li, Y.; Liang, Z.; He, X.; Li, X.; Tavajohi, N.; Li, B. A review of lithium-ion battery safety concerns: The issues, strategies, and testing standards. *Journal of Energy Chemistry*. **2021**, *59*, 83-99.
2. Liu, B.; Jia, Y.; Yuan, C.; Wang, L.; Gao, X.; Yin, S.; Xu, J. Safety issues and mechanisms of lithium-ion battery cell upon mechanical abusive loading: A review. *Energy Storage Materials*. **2020**, *24*, 85-112.
3. Shang, Z.; Yu, W.; Zhou, J.; Zhou, X.; Zeng, Z.; Tursun, R.; Liu, X.; Xu, S. Recycling of spent lithium-ion batteries in view of graphite recovery: A review. *eTransportation*. **2024**, *20*, 100320.
4. Yang, L.; Gao, Z.; Liu, T.; Huang, M.; Liu, G.; Feng, Y.; Shao, P.; Luo, X. Direct Electrochemical Leaching Method for High-Purity Lithium Recovery from Spent Lithium Batteries. *Environmental Science & Technology*. **2023**, *57*, 4591-4597.
5. Swain, B. Recovery and recycling of lithium: A review. *Separation and Purification Technology*. **2017**, *172*, 388-403.
6. Wang, J.; Koenig Jr, G. M. Direct Lithium Extraction Using Intercalation Materials. *Chemistry – A European Journal*. **2024**, *30*, e202302776.
7. Zhou, X.; Zhou, S.; Gao, Z.; Wang, G.; Zong, L.; Liu, J.; Zhu, F.; Ming, H.; Zheng, Y.; Chen, F.; Cao, N.; Yang, S. A statistical distribution-based pack-integrated model towards state estimation for lithium-ion batteries. *eTransportation*. **2024**, *19*, 100302.
8. Jin, P.; Wang, S.; Meng, Z.; Chen, B. China's lithium supply chains: Network evolution and resilience assessment. *Resources Policy*. **2023**, *87*, 104339.
9. Shao, L.; Jin, S. Resilience assessment of the lithium supply chain in China under impact of new energy vehicles and supply interruption. *Journal of Cleaner Production*. **2020**, *252*, 119624.
10. Meshram, P.; Pandey, B. D.; Mankhand, T. R. Extraction of lithium from primary and secondary sources by pre-treatment, leaching and separation: A comprehensive review. *Hydrometallurgy*. **2014**, *150*, 192-208.
11. Talens Peiró, L.; Villalba Méndez, G.; Ayres, R. U. Lithium: Sources, Production, Uses, and Recovery Outlook. *JOM*. **2013**, *65*, 986-996.
12. Yelatontsev, D.; Mukhachev, A. Processing of lithium ores: Industrial technologies and case studies – A review. *Hydrometallurgy*. **2021**, *201*, 105578.
13. Gu, G.; Gao, T. Sustainable production of lithium salts extraction from ores in China: Cleaner production assessment. *Resources Policy*. **2021**, *74*, 102261.
14. Vieceli, N.; Casasola, R.; Lombardo, G.; Ebin, B.; Petranikova, M. Hydrometallurgical recycling of EV lithium-ion batteries: Effects of incineration on the leaching efficiency of metals using sulfuric acid. *Waste Management*. **2021**, *125*, 192-203.
15. Fu, Y.; He, Y.; Li, J.; Qu, L.; Yang, Y.; Guo, X.; Xie, W. Improved hydrometallurgical extraction of valuable metals from spent lithium-ion batteries via a closed-loop process. *Journal of Alloys and Compounds*. **2020**, *847*, 156489.
16. Yang, S.; Zhang, F.; Ding, H.; He, P.; Zhou, H. Lithium Metal Extraction from Seawater. *Joule*. **2018**, *2*, 1648-1651.
17. Wang, J.; Yue, X.; Wang, P.; Yu, T.; Du, X.; Hao, X.; Abudula, A.; Guan, G. Electrochemical technologies for lithium recovery from liquid resources: A review. *Renewable and Sustainable Energy Reviews*. **2022**, *154*, 111813.
18. Farahbakhsh, J.; Arshadi, F.; Mofidi, Z.; Mohseni-Dargah, M.; Kök, C.; Assefi, M.; Soozanipour, A.; Zargar, M.; Asadnia, M.; Boroumand, Y.; Presser, V.; Razmjou, A. Direct lithium extraction: A new paradigm for lithium production and resource utilization. *Desalination*. **2024**, *575*, 117249.
19. Delmas, C.; Maccario, M.; Croguennec, L.; Le Cras, F.; Weill, F. Lithium deintercalation in LiFePO₄ nanoparticles via a domino-cascade model. *Nature Materials*. **2008**, *7*, 665-671.
20. Odetallah, M.; Kuss, C. A Review of Chemically Induced Intercalation and Deintercalation in Battery Materials. *Energy Technology*. **2023**, *11*, 2201060.
21. Allen, J. L.; Jow, T. R.; Wolfenstine, J. Kinetic Study of the Electrochemical FePO₄ to LiFePO₄ Phase

- Transition. *Chemistry of Materials*. **2007**, *19*, 2108-2111.
22. Zhao, X.; Yang, H.; Wang, Y.; Sha, Z. Review on the electrochemical extraction of lithium from seawater/brine. *Journal of Electroanalytical Chemistry*. **2019**, *850*, 113389.
 23. Battistel, A.; Palagonia, M. S.; Brogioli, D.; La Mantia, F.; Trócoli, R. Electrochemical Methods for Lithium Recovery: A Comprehensive and Critical Review. **2020**, *32*, e1905440.
 24. Patil, A.; Patil, V.; Wook Shin, D.; Choi, J.-W.; Paik, D.-S.; Yoon, S.-J. Issue and challenges facing rechargeable thin film lithium batteries. *Materials Research Bulletin*. **2008**, *43*, 1913-1942.
 25. Tarascon, J. M.; Armand, M. Issues and challenges facing rechargeable lithium batteries. *Nature*. **2001**, *414*, 359-67.
 26. He, L.; Xu, W.; Song, Y.; Luo, Y.; Liu, X.; Zhao, Z. New Insights into the Application of Lithium-Ion Battery Materials: Selective Extraction of Lithium from Brines via a Rocking-Chair Lithium-Ion Battery System. *Global Challenges*. **2018**, *2*, 1700079.
 27. Zhao, Z.; Si, X.; Liu, X.; He, L.; Liang, X. Li extraction from high Mg/Li ratio brine with LiFePO₄/FePO₄ as electrode materials. *Hydrometallurgy*. **2013**, *133*, 75-83.
 28. Zhao, Z.-w.; Si, X.-f.; Liang, X.-x.; Liu, X.-h.; He, L.-h. Electrochemical behavior of Li⁺, Mg²⁺, Na⁺ and K⁺ in LiFePO₄/FePO₄ structures. *Transactions of Nonferrous Metals Society of China*. **2013**, *23*, 1157-1164.
 29. Li, Z.; Liu, D.; Xiong, J.; He, L.; Zhao, Z.; Wang, D. Selective recovery of lithium and iron phosphate/carbon from spent lithium iron phosphate cathode material by anionic membrane slurry electrolysis. *Waste Management*. **2020**, *107*, 1-8.
 30. Sun, S.; Yu, X.; Li, M.; Duo, J.; Guo, Y.; Deng, T. Green recovery of lithium from geothermal water based on a novel lithium iron phosphate electrochemical technique. *Journal of Cleaner Production*. **2020**, *247*, 119178.
 31. Sata, T.; Mine, K.; Higa, M. Change in permselectivity between sulfate and chloride ions through anion exchange membrane with hydrophilicity of the membrane. *Journal of Membrane Science*. **1998**, *141*, 137-144.
 32. Mehanna, M.; Saito, T.; Yan, J.; Hickner, M.; Cao, X.; Huang, X.; Logan, B. E. Using microbial desalination cells to reduce water salinity prior to reverse osmosis. *Energy & Environmental Science*. **2010**, *3*, 1114-1120.
 33. Luo, G.; Li, X.; Chen, L.; Chao, Y.; Zhu, W. Electrochemical lithium ion pumps for lithium recovery: A systematic review and influencing factors analysis. *Desalination*. **2023**, *548*, 116228.
 34. Zhou, G.; Chen, L.; Chao, Y.; Li, X.; Luo, G.; Zhu, W. Progress in electrochemical lithium ion pumping for lithium recovery. *Journal of Energy Chemistry*. **2021**, *59*, 431-445.
 35. Luo, G.; Zhu, L.; Li, X.; Zhou, G.; Sun, J.; Chen, L.; Chao, Y.; Jiang, L.; Zhu, W. Electrochemical lithium ions pump for lithium recovery from brine by using a surface stability Al₂O₃-ZrO₂ coated LiMn₂O₄ electrode. *Journal of Energy Chemistry*. **2022**, *69*, 244-252.
 36. Battistel, A.; Palagonia, M. S.; Brogioli, D.; La Mantia, F.; Trócoli, R. Electrochemical Methods for Lithium Recovery: A Comprehensive and Critical Review. *Advanced materials*. **2020**, *32*, 1905440.
 37. Calvo, E. J. Electrochemical methods for sustainable recovery of lithium from natural brines and battery recycling. *Current Opinion in Electrochemistry*. **2019**, *15*, 102-108.
 38. Pasta, M.; Wessells, C. D.; Cui, Y.; La Mantia, F. A Desalination Battery. *Nano letters*. **2012**, *12*, 839-843.
 39. Han, T.; Yu, X.; Guo, Y.; Li, M.; Duo, J.; Deng, T. Green recovery of low concentration of lithium from geothermal water by a novel FPO/KNiFC ion pump technique. *Electrochimica Acta*. **2020**, *350*, 136385.
 40. Sasaki, K.; Shin-mura, K.; Honda, S.; Tazoe, H.; Niwa, E. A three-electrode dual-power-supply electrochemical pumping system for fast and energy efficient lithium extraction and recovery from solutions. *Communications Engineering*. **2024**, *3*, 29.
 41. Li, Z.; Li, C.; Liu, X.; Cao, L.; Li, P.; Wei, R.; Li, X.; Guo, D.; Huang, K.-W.; Lai, Z. Continuous electrical pumping membrane process for seawater lithium mining. *Energy & Environmental Science*. **2021**, *14*, 3152-3159.
 42. Al-Amshawee, S.; Yunus, M. Y. B. M.; Azoddein, A. A. M.; Hassell, D. G.; Dakhil, I. H.; Hasan, H. A. Electrodialysis desalination for water and wastewater: A review. *Chemical Engineering Journal*. **2020**, *380*, 122231.
 43. Gonçalves, F.; Fernandes, C.; Cameira dos Santos, P.; de Pinho, M. N. Wine tartaric stabilization by electrodialysis and its assessment by the saturation temperature. *Journal of Food Engineering*. **2003**, *59*, 229-235.
 44. Zhao, D.; Lee, L. Y.; Ong, S. L.; Chowdhury, P.; Siah, K. B.; Ng, H. Y. Electrodialysis reversal for industrial reverse osmosis brine treatment. *Separation and Purification Technology*. **2019**, *213*, 339-347.
 45. Gmar, S.; Chagnes, A. Recent advances on electrodialysis for the recovery of lithium from primary and

- secondary resources. *Hydrometallurgy*. **2019**, *189*, 105124.
46. Song, Y.; Zhao, Z. Recovery of lithium from spent lithium-ion batteries using precipitation and electro dialysis techniques. *Separation and Purification Technology*. **2018**, *206*, 335-342.
 47. Wang, Q.; Du, X.; Gao, F.; Liu, F.; Liu, M.; Hao, X.; Tang, K.; Guan, G.; Abudula, A. A novel $\text{H}_{1.6}\text{Mn}_{1.6}\text{O}_4$ /reduced graphene oxide composite film for selective electrochemical capturing lithium ions with low concentration. *Separation and Purification Technology*. **2019**, *226*, 59-67.
 48. Wang, W.; Hong, G.; Zhang, Y.; Yang, X.; Hu, N.; Zhang, J.; Sorokin, P.; Shao, L. Designing an energy-efficient multi-stage selective electro dialysis process based on high-performance materials for lithium extraction. *Journal of Membrane Science*. **2023**, *675*, 121534.
 49. Li, B.; Jiang, L.; Xiao, N.; Liu, S.; Zhang, Z.; Liu, F.; Free, M. L. Enhanced lithium separation with $\text{Li}_{1.3}\text{Al}_{0.3}\text{Ti}_{1.7}(\text{PO}_4)_3$ lithium superionic conductor and aided charge balance. *Separation and Purification Technology*. **2024**, *351*, 128058.
 50. Wang, D.; Zhang, Y.; Pang, B.; Liu, C.; Liu, Z. Study on the optimal conductivity titration parameters for SO_4^{2-} in cement-based materials. *Measurement*. **2024**, 115277.
 51. Wenten, I. G.; Bazant, M. Z.; Khoiruddin, K. Mitigating electro dialysis membrane fouling in seawater desalination. *Separation and Purification Technology*. **2024**, *345*, 127228.

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